

* %yield could be in **g** or **mol**

$$\% \text{yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

$$= \frac{7.54\text{g CaO}}{8.97\text{g CaO}} \times 100 = 0.841 \times 100$$
$$= \boxed{84.1 \%}$$

Percent Yield Holt Chemfile Answers

Clifford Lane



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Holt Chemistry Ralph Thomas Myers,2004 **Holt Chemistry** ,2003 Holt ChemFile Lab Program ,2005

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