Chapter 4 - Practice Problems Solutions

1) Complete and balance the following reactions. Remember to form the stable product and then balance. Assume the product is solid. What type of reactions are these?

_____ Mg (s) + _____ O_2 (g) →

2 Mg (s) + O_2 (g) → 2 MgO (s)

_____ Na (s) + _____ Cl_2 (g) →

2 Na (s) + Cl_2 (g) → 2 NaCl (s)

_____ Al (s) + _____ O_2 (g) →

4 Al (s) + 3 O_2 (g) → 2 Al_2O_3 (s)

_____ Ca (s) + _____ Br_2 (g) →

Ca (s) + Br_2 (g) → $CaBr_2$ (s)

These are all combination reactions.

2) Balance the following reactions. What type of reactions are these?

_____C_3H_8 (g) + _____O_2 (g) → _____CO_2 (g) + _____ H_2O (g)

C_3H_8 (g) + 5 O_2 (g) → 3 CO_2 (g) + 4 H_2O (g)

_____C_5H_6 (g) + _____O_2 (g) → _____CO_2 (g) + _____ H_2O (g)

2 C_6H_6 (g) + 15 O_2 (g) → 12 CO_2 (g) + 6 H_2O (g)

These are both combustion reactions.

3) Balance the following reactions. What type of reactions are these?

_____ H_2O (l) → _____H_2 (g) + _____O_2 (g)

2 H_2O (l) → 2 H_2 (g) + O_2 (g)

_____ $KClO_3$ (s) → _____KCl (s) + _____O_2 (g)

2 $KClO_3$ (s) → 2 KCl (s) + 3 O_2 (g)

These are both decomposition reactions.

4) Balance the following reactions. What type of reactions are these?

_____ Mg (s) + _____HCl (aq) →_____$MgCl_2$ (aq) + _____H_2 (g)

Mg (s) + 2 HCl (aq) → $MgCl_2$ (aq) + H_2 (g)

_____ Ag (s) + _____H_2S (aq) →_____Ag_2S (s) + _____H_2 (g)

2 Ag (s) + H_2S (aq) → Ag_2S (s) + H_2 (g)

These are both single displacement reactions.

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Connie Lockhart Ellefson

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