

## Reaction Rates Test Name:

20 K 9 T 9 C 8 A

## Assessment of Learning (Product)

Multiple Choice; Circle the most correct answer. (20K)

1. The kinetics of the reaction below were studied and it was determined that the reaction rate increased by a factor of 9 when the concentration of B was tripled. The reaction is \_\_\_\_\_ order in B.



- A) zero      B) first      C) second      D) third      E) one-half

2. Consider the following reaction:



If the rate of formation of  $NH_3$  is  $4.0 \times 10^{-4}$  mol/s, then the rate of consumption of  $H_2$  is:

- A)  $2.0 \times 10^{-4}$  mol/s    B)  $4.0 \times 10^{-4}$  mol/s    C)  $6.0 \times 10^{-4}$  mol/s    D)  $1.2 \times 10^{-3}$  mol/s

3. Which of the following could be used to describe the rate of a reaction?

- A)  $\frac{\text{change in time}}{\text{change in concentration}}$       B)  $\frac{\text{change in mass}}{\text{change in concentration}}$   
C)  $\frac{\text{change in concentration}}{\text{change in time}}$       D)  $\frac{\text{change in concentration}}{\text{change in mass}}$

4. Which of the following would **not** be units for reaction rate?

- A) g/s      B) M/min      C) kJ/mol      D) mol/min

5. Consider the following reaction:



The activation energy of the forward reaction is 134 kJ. What is the activation energy for the reverse reaction?

- A) -134 kJ      B) -100 kJ      C) 234 kJ      D) 368 kJ

6. Which of the following changes will increase the average kinetic energy of reactant molecules?

- A) adding a catalyst      B) increasing the temperature  
C) increasing the surface area      D) increasing the concentration

7. The minimum amount of energy needed to start a reaction is called the:

- A) activation energy      B) energy of reaction  
C) entropy of reaction      D) reaction mechanism energy

8. In general, a chemical reaction requiring a large activation energy will proceed:

- A) at a fast rate      B) at a slow rate  
C) only at low temperatures      D) only at low concentrations

# Reaction Rates Assessment Answers

**SJ Ball**



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