

THERMOCHEMISTRY ASSIGNMENT

1. Thermodynamic stability of graphite is defined as the heat required for conversion of graphite into diamond. The enthalpy change for this reaction is, 1.9 kJ. Sublimation enthalpy of graphite is 716.68 kJ/mol. The enthalpy of melting of diamond is 1.9 kJ/mol. Calculate the enthalpy of fusion of graphite.

The above reaction is exothermic and as such will release energy and decrease the temperature of the hot water the graphite will be being produced. As a result the hot water will be cooled down and the graphite will get into liquid state.

2. Thermodynamic heat capacity of diamond is 1.9 kJ/mol. The enthalpy of fusion of diamond is 1.9 kJ/mol. Calculate the enthalpy of fusion of diamond.

$$\Delta H_{\text{fusion}} = \Delta H_{\text{sublimation}} - \Delta H_{\text{melting}}$$

$$\Delta H_{\text{fusion}} = 1.9 \text{ kJ/mol} - 1.9 \text{ kJ/mol} = 0 \text{ kJ/mol}$$

$$\Delta H_{\text{fusion}} = 0 \text{ kJ/mol}$$

$$\Delta H_{\text{fusion}} = 0 \text{ kJ/mol}$$

3. Thermodynamic heat capacity of graphite is 1.9 kJ/mol. The enthalpy of fusion of graphite is 1.9 kJ/mol. Calculate the enthalpy of fusion of graphite.

$$\Delta H_{\text{fusion}} = \Delta H_{\text{sublimation}} - \Delta H_{\text{melting}}$$

$$\Delta H_{\text{fusion}} = 1.9 \text{ kJ/mol} - 1.9 \text{ kJ/mol}$$

$$\Delta H_{\text{fusion}} = 0 \text{ kJ/mol}$$

$$\Delta H_{\text{fusion}} = 0 \text{ kJ/mol}$$

4. Calculate the enthalpy of fusion of graphite from 1.9 kJ/mol and 1.9 kJ/mol.

$$\Delta H_{\text{fusion}} = \Delta H_{\text{sublimation}} - \Delta H_{\text{melting}}$$

$$\Delta H_{\text{fusion}} = 1.9 \text{ kJ/mol} - 1.9 \text{ kJ/mol}$$

$$\Delta H_{\text{fusion}} = 0 \text{ kJ/mol}$$

$$\Delta H_{\text{fusion}} = 1.9 \text{ kJ/mol} - 1.9 \text{ kJ/mol}$$

$$\Delta H_{\text{fusion}} = 0 \text{ kJ/mol}$$

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5. Calculate the enthalpy of fusion of graphite from 1.9 kJ/mol and 1.9 kJ/mol. The enthalpy of fusion of graphite is 1.9 kJ/mol.

$$\Delta H_{\text{fusion}} = \Delta H_{\text{sublimation}} - \Delta H_{\text{melting}}$$

$$\Delta H_{\text{fusion}} = 1.9 \text{ kJ/mol} - 1.9 \text{ kJ/mol}$$

$$\Delta H_{\text{fusion}} = 0 \text{ kJ/mol}$$

$$\Delta H_{\text{fusion}} = 0 \text{ kJ/mol}$$

$$\Delta H_{\text{fusion}} = 0 \text{ kJ/mol}$$

$$\Delta H_{\text{fusion}} = 0 \text{ kJ/mol}$$

6. Calculate the enthalpy of fusion of graphite from 1.9 kJ/mol and 1.9 kJ/mol. The enthalpy of fusion of graphite is 1.9 kJ/mol.

Graphite is a solid at 25°C. It is added to the water and the temperature is 25°C. The enthalpy of fusion of graphite is 1.9 kJ/mol. The enthalpy of fusion of graphite is 1.9 kJ/mol.

Reaction Rates And Thermochemistry Answers

Guillaume Favre



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