

## Chapter 5 • Periodicity Activity

## ATOMIC RADII

**Directions:** Using the given data for the sizes of atoms (atomic radii) below,

1. Create a line graph (use a ruler) by graphing **Period 2's** data on the graph below.
2. In a **different color**, create a line graph **Period 3's** data on the same graph.
3. Label each axis, and include a key/legend.
4. Answer all questions.

	1	2	13	14	15	16	17	18
1	H 1	<b>Atomic Radii</b>						He 2
2	Li 123	Be 89	B 88	C 77	N 70	O 66	F 64	Ne —
3	Na 157	Mg 136	Al 125	Si 117	P 110	S 104	Cl 99	Ar —
4	K 203	Ca 174	Ga 125	Ge 122	As 121	Se 117	Br 114	Kr —
5	Rb 216	Sr 192	In 150	Sn 140	Sb 141	Te 137	I 133	Xe —
6	Cs 234	Ba 198	Tl 155	Pb 154	Bi 152	Po 153	At —	Rn —

**Note:** Mark the vertical axis from 0 – 160 (by 20's)

### Trend in Atomic Size Across the Periodic Table

Li	Be	B	C	N	O	F
Na	Mg	Al	Si	P	S	Cl

### Horizontal Trend Questions

1. What is the trend in atomic size as you go ~~across~~ the periodic table? \_\_\_\_\_
2. This is true because there are more \_\_\_\_\_ (protons / electrons / layers of electrons) pulling the electron cloud in toward the nucleus.

### Vertical Trend Questions

4. This is true because there are more \_\_\_\_\_ (protons / electrons / layers of electrons) making the electron cloud larger.

# Periodicity Activity Answers

**Tobias Bleicker**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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