ANSWER KEY

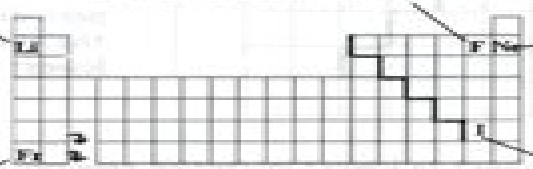
Trends on the Periodic Table

Part 1: Identifying chemical properties/characteristics of the elements

For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)

Lithium
Being one of the smaller elements with only 3 electrons, lithium is happy to give one of its electrons away and never seems to ask for any in return.

Fluorine
One of the smallest elements, Fluorine is often described as scrappy with a "magnetic" personality. As part of the halogen family (notorious electron thieves), Fluorine is the best at convincing other elements to give up their electrons to him. However, he holds on tight to his own electrons and will not let them go without a fight.



Neon
Part of the Noble family, Neon does not readily interact with other elements. With a full set of valence electrons, Neon feels that he has everything he needs to make him happy without sharing with others.

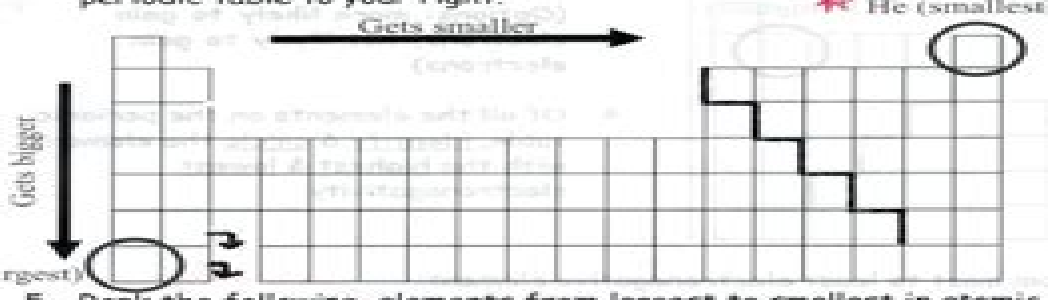
Francium
Francium is the largest of all of the elements and has so many electrons that it has trouble keeping track of them. Francium will eagerly give up electrons to other elements and never takes in extras. But watch out, it has an extremely reactive and explosive personality!

Iodine
As part of the halogen family, Iodine is also a skilled electron thief, however, not nearly as good as its smaller brothers and sisters. Iodine will also put up a pretty good fight to keep its own electrons.

Part 2: Use your textbook to complete the following:

Atomic Radii Size:

1. Explain **atomic radius**: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.
2. Using the information in figure 5-13 (pg 141) in your book, **write** the atomic radii size for **alkali metals** & **period #2** on the periodic table to your right.



3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. **(Options: gets bigger OR gets smaller)**
4. Of all the elements on the periodic table, identify & circle the largest & smallest in size.
5. Rank the following elements from largest to smallest in atomic radii:
   a) Be, Ba, Mg = Ba, Mg, Be
   b) Li, F, B = Li, B, F
6. How can you explain the trend:
   a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on electrons, bring the electron cloud closer. Effective nuclear charge, Zeff increases the attraction of the nucleus and therefore pulls the electron cloud closer to the nucleus resulting in smaller radius

not a good enough answer

AP Answer

Periodic Trends Pogil Answers Extension Questions

Scott C. Dulebohn

Periodic Trends Pogil Answers Extension Questions:

Chemistry and Our Universe , Return to the periodic table introduced in Lecture 1 to practice predicting properties of elements based on their electronic structure Then witness what happens when three different alkali metals react with water Theory forecasts a pronounced difference in the result Is there Approximating Periodic and Non-periodic Trends in Time-series Data Carlotta Ching Ting Fok,2002 The new model is then applied to Brown and Moskowitz s time series data to investigate the long term evolution to the four interpersonal behaviors and to the GDP data to examine the periodic and non periodic pattern for the GDP values of the 16 countries Finally the extent to which the model is accurate is tested using simulated data *Chemical Periodicity* Robert Thomas Sanderson,2013-04-20

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