

Name _____

Period _____

Date Tuesday, November 1, 2016

Periodic Trends

ATOMIC RADIUS

1. What trend in atomic radius do you see as you go down a group/family on the periodic table?
2. What causes this trend?
3. What trend in atomic radius do you see as you go across a period/row on the periodic table?
4. What causes this trend?
5. Circle the atom in each pair that has the largest atomic radius.
 - a) Al B
 - b) S O
 - c) Br Cl
 - d) Na Al
 - e) O F
 - f) Mg Ca

6. Put the following elements in order from smallest to largest atomic radius **and** explain why:
C, O, Sn, Sr.

ELECTRONEGATIVITY

7. Define electronegativity
8. How does the ionic radius of a nonmetal compare with its atomic radius?
9. What trend in electronegativity do you see as you go down a group/family on the periodic table?
10. What causes this trend?
11. What trend in electronegativity do you see as you go across a period/row on the periodic table?
12. What causes this trend?
13. Circle the atom in each pair that has the greater electronegativity.
 - a) Ca Ga
 - b) Li O
 - c) Cl S
 - d) Br As
 - e) Ba Sr
 - f) O S

GENERAL QUESTIONS

14. Which group tends to form +1 ions? _____
15. Which group tends to form +2 ions? _____
16. Which group tends to form -1 ions? _____
17. Which group tends not to form ions or react? _____

Periodic Trends Exam

L Darling-Hammond



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Table of Contents **Periodic Trends Exam**

1. Understanding the eBook Periodic Trends Exam
 - The Rise of Digital Reading Periodic Trends Exam
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Exam
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Exam
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Exam
 - Personalized Recommendations
 - Periodic Trends Exam User Reviews and Ratings
 - Periodic Trends Exam and Bestseller Lists

5. Accessing Periodic Trends Exam Free and Paid eBooks
 - Periodic Trends Exam Public Domain eBooks
 - Periodic Trends Exam eBook Subscription Services
 - Periodic Trends Exam Budget-Friendly Options
6. Navigating Periodic Trends Exam eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Exam Compatibility with Devices
 - Periodic Trends Exam Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Exam
 - Highlighting and Note-Taking Periodic Trends Exam
 - Interactive Elements Periodic Trends Exam
8. Staying Engaged with Periodic Trends Exam
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Exam
9. Balancing eBooks and Physical Books Periodic Trends Exam
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Exam
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Trends Exam
 - Setting Reading Goals Periodic Trends Exam
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Trends Exam
 - Fact-Checking eBook Content of Periodic Trends Exam
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

-
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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