

Percent Composition and Molecular Formula Worksheet Key

1. What's the empirical formula of a molecule containing 65.5% carbon, 5.5% hydrogen, and 29.0% oxygen? **C_3H_3O mass = 55 g/mole**
2. If the molar mass of the compound in problem 1 is 110 grams/mole, what's the molecular formula? **$C_6H_6O_2$**
3. What's the empirical formula of a molecule containing 18.7% lithium, 16.3% carbon, and 65.0% oxygen? **Li_2CO_3**
4. If the molar mass of the compound in problem 3 is 73.8 grams/mole, what's the molecular formula? **Li_2CO_3**

Write the molecular formulas of the following compounds:

5. A compound with an empirical formula of C_2OH_4 and a molar mass of 88 grams per mole. **$C_4O_2H_8$**
6. A compound with an empirical formula of C_4H_4O and a molar mass of 136 grams per mole. **$C_8H_8O_2$**
7. A compound with an empirical formula of $CFBrO$ and a molar mass of 254.7 grams per mole. **$C_2F_2Br_2O_2$**
8. A compound with an empirical formula of C_2H_8N and a molar mass of 46 grams per mole. **C_2H_8N**

Answer the following questions:

9. The percentage composition of acetic acid is found to be 39.9% C, 6.7% H, and 53.4% O. Determine the empirical formula of acetic acid. **CH_2O**
10. The molar mass for question #9 was determined by experiment to be 60.0 g/mol. What is the molecular formula? **$C_2H_4O_2$**
11. Aniline, a starting material for urethane plastic foams, consists of C, H, and N. Combustion of such compounds yields CO_2 , H_2O , and N_2 as products. If the combustion of 9.71 g of aniline yields 6.63 g H_2O and 1.46 g N_2 , what is its empirical formula? **C_6H_7N**
12. The molar mass of aniline is 93 g/mol. What is its molecular formula? **C_6H_7N**
13. Calculate the mass percent of carbon, nitrogen and oxygen in acetamide, C_2H_5NO . **%C 40.668 %H 8.533 %N 23.713 %O 27.086**
14. A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms were produced. What is the empirical formula of the compound? **PCl_3**
15. When 2.5000 g of an oxide of mercury, (Hg_xO_y) is decomposed into the elements by heating, 2.405 g of mercury are produced. Calculate the empirical formula. **Hg_2O**
16. The compound benzamide has the following percent composition. What is the empirical formula?
C = 69.40 % H= 5.825 % O = 13.21 % N= 11.57 % **C_7H_7NO**
17. A component of protein called serine has an approximate molar mass of 100 g/mole. If the percent composition is as follows, what is the empirical and molecular formula of serine?
C = 34.95 % H= 6.844 % O = 46.56 % N= 13.59 %

$C_3H_7NO_3$ empirical formula

$C_3H_7NO_3$ molecular formula

Percent Composition And Molecular Formula Answers

K Morrison



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Applied Chemistry for Polytechnic and Engineering Courses Dr. R.S. Chauhan, 2021-05-27 The book includes the following chapters in details Language of Chemistry Atomic Structure The Periodic table and Atomic properties Water Chemical Bonding Solutions Electrolysis Environmental Chemistry Experiments **State Board Questions and Answers**, 1911 CliffsNotes Chemistry Practice Pack Charles Henrickson, 2010-02-08 About the Contents Pretest Helps you pinpoint where you need the most help Topic Area Reviews Measurement and Units of Measurement Matter Elements Compounds and Mixtures Atoms I The Basics Formulas and Names of Ionic Compounds Acids and Bases The Mole Elements and Compounds Percent Composition and Empirical and Molecular Formulas Chemical Reactions and Chemical Equations Calculations Using Balanced Equations Atoms II Atomic Structure and Periodic Properties Chemical Bonding The Formation of Compounds Gases and the Gas Laws The Forces between Molecules Solids and Liquids Solutions and Solution Composition Acids Bases and Neutralization Glossary Customized Full Length Exam Covers all subject areas Pretest that pinpoints what you need to study most Clear concise reviews of every topic Targeted example problems in every chapter with solutions and explanations Customized full length exam that adapts to your skill level **Chemistry** Richard Post, Chad

Snyder, Clifford C. Houk, 2020-08-17 A practical complete and easy to use guide for understanding major chemistry concepts and terms Master the fundamentals of chemistry with this fast and easy guide Chemistry is a fundamental science that touches all other sciences including biology physics electronics environmental studies astronomy and more Thousands of students have successfully used the previous editions of Chemistry Concepts and Problems A Self Teaching Guide to learn chemistry either independently as a refresher or in parallel with a college chemistry course This newly revised edition includes updates and additions to improve your success in learning chemistry This book uses an interactive self teaching method including frequent questions and study problems increasing both the speed of learning and retention Monitor your progress with self tests and master chemistry quickly This revised Third Edition provides a fresh step by step approach to learning that requires no prerequisites lets you work at your own pace and reinforces what you learn ensuring lifelong mastery Master the science of basic chemistry with this innovative self paced study guide Teach yourself chemistry refresh your knowledge in preparation for medical studies or other coursework or enhance your college chemistry course Use self study features including review questions and quizzes to ensure that you re really learning the material Prepare for a career in the sciences medicine or engineering with the core content in this user friendly guide Authored by expert postsecondary educators this unique book gently leads students to deeper levels and concepts with practice critical thinking problem solving and self assessment at every stage *NEET National Eligibility Cum Entrance Test Chemistry Class 11 Volume I*

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 - Chapter 1: Introduction to Percent Composition And Molecular Formula Answers
 - Chapter 2: Essential Elements of Percent Composition And Molecular Formula Answers
 - Chapter 3: Percent Composition And Molecular Formula Answers in Everyday Life
 - Chapter 4: Percent Composition And Molecular Formula Answers in Specific Contexts
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- The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Percent Composition And Molecular Formula Answers.

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