

# of substance:

## Substance:

Mass of particles of a substance  
in mol

6.02 x 10<sup>23</sup> particles  
amount of substance

## Constant:

particles per mol, 6.02 x 10<sup>23</sup> mol<sup>-1</sup>

## Unit:

mol, measured in g  
equivalent to Mr

## Atomic Structure

### Isotopes:

→ Same # of protons, same element  
of p<sup>+</sup> and e<sup>-</sup> but different  
neutrons different mass

### Physical Properties:

→ Can be different - e.g. different  
melting + boiling points  
→ i.e. heavy water has higher melting  
point + boiling point + density

### Chemical Properties:

→ Same no. of electrons + same electronic  
structure in outer shell + transfer of e<sup>-</sup> is  
unaffected hence have same chemical  
properties.  
→ Different no. of n has no effect  
on bonding + reactions

### Compounds + ions FORMULAE:

Alkali: NO<sub>3</sub><sup>-</sup>

### Molar Gas Volume:

→ Volume of gas per mole of gas molecules in a stated temperature + pressure  
- Measured in dm<sup>3</sup> mol<sup>-1</sup> or m<sup>3</sup> mol<sup>-1</sup>

At RTP (20°C and 100kPa):  
(293K and 100000 Pa)

is 24 dm<sup>3</sup> mol<sup>-1</sup>  
at RTP.

Volume /  
[dm<sup>3</sup>]

## Electronegativity:

→ the measure of attraction of a bonded atom  
in a covalent bond

### ELECTRONEGATIVITY INCREASES ACROSS PERIOD:

- Increase in no. of protons in an atom  
→ Nuclear charge increases

### ELECTRONEGATIVITY DE:

- More electron shells  
→ More shielding  
→ Atomic radius increases  
→ Nuclear charge decreases  
→ Nuclear attraction decreases  
→ Electronegativity decreases

# OCR A LEVEL CHEMISTRY MODULE 2 DIAGRAMS



→ Consists of a small dense  
nucleus of protons + neutrons  
surrounded by electrons orbiting  
in electron shells in shells

- No. of p<sup>+</sup> = No. of e<sup>-</sup>

→ Equal + opposite charges cancel  
in atom has neutral charge  
- Neutrons important in nuclear  
stability - Electrostatic repulsion  
between p<sup>+</sup> in small dense nucleus

### Mass Number:

→ No. of p<sup>+</sup> + n in nucleus of atom  
- p<sup>+</sup> and n have relative mass of 1  
and e<sup>-</sup> mass is negligible is 0

### Relative Isotopic Mass:

→ The mass of an atom of an isotope  
relative to 1/12th the mass of an atom of C-12

### Relative Atomic Mass:

→ The weighted mean mass of an atom of  
an element relative to 1/12th the mass of  
an atom of C-12

### Weighted Mean Mass:

→ Average mass of all isotopes, taking into  
account the relative abundance of each isotope

- It dipoles  
Molecule is  
out + inlets

e.g. H<sub>2</sub>O

- O is more electronegative than H  
- O-H has polar bond  
since do not cancel there is an overall dipole

e.g. CO<sub>2</sub>

O is more electronegative than C  
C=O has a dipole  
but the bond is symmetrical so the dipoles cancel

# Ocr Chemistry A Level Unit 2 2014 Paper

**Mike Smith**



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  - The Rise of Digital Reading Ocr Chemistry A Level Unit 2 2014 Paper
  - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry A Level Unit 2 2014 Paper
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ocr Chemistry A Level Unit 2 2014 Paper
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry A Level Unit 2 2014 Paper
  - Personalized Recommendations
  - Ocr Chemistry A Level Unit 2 2014 Paper User Reviews and Ratings
  - Ocr Chemistry A Level Unit 2 2014 Paper and Bestseller Lists
5. Accessing Ocr Chemistry A Level Unit 2 2014 Paper Free and Paid eBooks

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- Ocr Chemistry A Level Unit 2 2014 Paper eBook Subscription Services
- Ocr Chemistry A Level Unit 2 2014 Paper Budget-Friendly Options
- 6. Navigating Ocr Chemistry A Level Unit 2 2014 Paper eBook Formats
  - ePub, PDF, MOBI, and More
  - Ocr Chemistry A Level Unit 2 2014 Paper Compatibility with Devices
  - Ocr Chemistry A Level Unit 2 2014 Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ocr Chemistry A Level Unit 2 2014 Paper
  - Highlighting and Note-Taking Ocr Chemistry A Level Unit 2 2014 Paper
  - Interactive Elements Ocr Chemistry A Level Unit 2 2014 Paper
- 8. Staying Engaged with Ocr Chemistry A Level Unit 2 2014 Paper
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ocr Chemistry A Level Unit 2 2014 Paper
- 9. Balancing eBooks and Physical Books Ocr Chemistry A Level Unit 2 2014 Paper
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ocr Chemistry A Level Unit 2 2014 Paper
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ocr Chemistry A Level Unit 2 2014 Paper
  - Setting Reading Goals Ocr Chemistry A Level Unit 2 2014 Paper
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ocr Chemistry A Level Unit 2 2014 Paper
  - Fact-Checking eBook Content of Ocr Chemistry A Level Unit 2 2014 Paper
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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