

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance Basic SI Unit? kg
 - Volume of an object? ruler or water displacement method Basic SI Unit? cm³ or ml
 - Length of an object? Ruler Basic SI Unit? meter
 - Density of an object? graduated cylinder Basic SI Unit? ml
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 25 g and has a volume of 3 cm³? 8g/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.
solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of neutrons and protons.

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Table of Contents Science Second Semester Final Exam Study Guide

1. Understanding the eBook Science Second Semester Final Exam Study Guide
 - The Rise of Digital Reading Science Second Semester Final Exam Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Science Second Semester Final Exam Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Science Second Semester Final Exam Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Science Second Semester Final Exam Study Guide
 - Personalized Recommendations
 - Science Second Semester Final Exam Study Guide User Reviews and Ratings
 - Science Second Semester Final Exam Study Guide and Bestseller Lists
5. Accessing Science Second Semester Final Exam Study Guide Free and Paid eBooks
 - Science Second Semester Final Exam Study Guide Public Domain eBooks
 - Science Second Semester Final Exam Study Guide eBook Subscription Services
 - Science Second Semester Final Exam Study Guide Budget-Friendly Options
6. Navigating Science Second Semester Final Exam Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Science Second Semester Final Exam Study Guide Compatibility with Devices
 - Science Second Semester Final Exam Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Science Second Semester Final Exam Study Guide
 - Highlighting and Note-Taking Science Second Semester Final Exam Study Guide
 - Interactive Elements Science Second Semester Final Exam Study Guide
8. Staying Engaged with Science Second Semester Final Exam Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Science Second Semester Final Exam Study Guide
- 9. Balancing eBooks and Physical Books Science Second Semester Final Exam Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Science Second Semester Final Exam Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Science Second Semester Final Exam Study Guide
 - Setting Reading Goals Science Second Semester Final Exam Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Science Second Semester Final Exam Study Guide
 - Fact-Checking eBook Content of Science Second Semester Final Exam Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
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 - Interactive and Gamified eBooks

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