

## Reaction Rates and Potential Energy Diagrams

1. Chemical reactions occur when reactants collide. For what reasons may a collision fail to produce a chemical reaction?

Not enough energy; improper angle.

2. If every collision between reactants leads to a reaction, what determines the rate at which the reaction occurs?

Nature of reactants, Concentration, Temperature, Catalysts.

3. What is the activation energy of a reaction, and how is this energy related to the activated complex of the reaction?

$E_a$  is the minimum amount of energy for a reaction to occur. It is the amount of energy required to create an activated complex.

4. What happens when a catalyst is used in a reaction?

A catalyst changes the reaction mechanism, in the process lowering the activation energy.

5. Name 4 things that will speed up or slow down a chemical reaction.

Increase concentration by distillation of a solvent, Increase concentration by increasing pressure of a gas, Increase temp, Add a catalyst, Add an inhibitor.

6. Draw an energy diagram for a reaction. Label the axis, PE of reactants = 350 KJ/mol,  $E_a$  = 100 KJ/mol, PE of products = 250 KJ/mol.



7. Is the reaction in # 6 exothermic or endothermic? Explain.

Exothermic. The  $\Delta H$  is -100 KJ/mol which means heat is released.

8. How could you lower the activation energy for the reaction in #6?

Add a catalyst.

# Potential Energy Diagram Practice Problems

**R Sanford**



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