

Configuration du serveur du système Polycom® CMA™

Ce chapitre décrit les différentes opérations à effectuer pour configurer le serveur d'un système Polycom® Converged Management Application™ (CMA™) et le connecter au réseau.



Si vous installez une configuration redondante de Polycom CMA, exécutez toutes les procédures figurant dans ce chapitre sur les deux serveurs, sauf instruction contraire.

Réunir les éléments nécessaires

Avant d'installer un système Polycom CMA, vous devez réunir les différents éléments suivants :

- *Notes de diffusion du système Polycom CMA*
- Serveur du système Polycom CMA
- Étude complète du site ou boîte à outils du projet
- Ordinateur comportant un port série ou un port Ethernet

Si vous souhaitez vous connecter au serveur du système Polycom CMA par un port Ethernet, vous devez disposer également des composants suivants :

- Un câble croisé OU
- Un concentrateur ou un commutateur et deux câbles Ethernet

Polycom Cma 40user Guide

Chao Zhang



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web jan 26 2016 the solubility product or ksp is a way of showing how effective a chemical is at dissolving into an aqueous solution the higher the ksp the more the chemical will dissolve into that solvent for this lab calcium hydroxide is created then ionized Ca(OH)_2 when calcium is put in water reacts as follows $\text{Ca(OH)}_2 \rightleftharpoons \text{Ca}^{2+} + 2\text{OH}^-$

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web k sp of calcium hydroxide 1 author j m mccormick last update october 13 2013 introduction an equilibrium constant K is related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined with eqn 2 to determine ΔH and ΔS from the temperature dependence of K

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web in today s experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the

concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$ the solubility product expression describes in mathematical terms the equilibrium

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web the chemical equation used was $\text{Ca(OH)}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$ the solubility product was determined to be ksp $\text{Ca}^{2+}\text{OH}^{-}$ from the titration which was made by dripping a concentrated solution of calcium hydroxide into hydrochloric acid the calcium hydroxide was filtered before starting the titration

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web jul 12 2023 because the concentration of a pure solid such as Ca_3PO_4 is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $\text{Ca}_3\text{PO}_4(\text{s}) \rightleftharpoons 3\text{Ca}^{2+}(\text{aq}) + 4\text{PO}_4^{3-}(\text{aq})$ $K_{\text{sp}} = [\text{Ca}^{2+}]^3 [\text{PO}_4^{3-}]^4$

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accepted value for K_{sp} for sodium hydroxide is 5×10^{-6}