

Unit 3 Notes: Periodic Table Notes

- John Newlands proposed an organization system based on increasing atomic mass in 1864.
- He noticed that both the chemical and physical properties repeated every 8 elements and called this the Law of Octaves.
- In 1869 both Lothar Meyer and Dmitri Mendeleev showed a connection between atomic mass and an element's properties.
- Mendeleev published first, and is given credit for this.
- He also noticed a periodic pattern when elements were ordered by increasing Atomic Mass.
- By arranging elements in order of increasing atomic mass into columns, Mendeleev created the first Periodic Table.
- This table also predicted the existence and properties of undiscovered elements.
- After many new elements were discovered, it appeared that a number of elements were out of order based on their Properties.
- In 1913 Henry Mosley discovered that each element contains a unique number of Protons.
- By rearranging the elements based on Atomic Number, the problems with the Periodic Table were corrected.
- This new arrangement creates a periodic repetition of both physical and chemical properties known as the Periodic Law.

PERIODS

Periods are the Rows

Valence electrons across a period are in the same energy level

GROUPS

Groups/Families are the Columns

There are equal numbers of valence electrons in a group.

- When elements are arranged in order of increasing Atomic Number, there is a periodic repetition of their physical and chemical properties

Periodic Trends Packet Key

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Table of Contents Periodic Trends Packet Key

1. Understanding the eBook Periodic Trends Packet Key
 - The Rise of Digital Reading Periodic Trends Packet Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Packet Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Packet Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Packet Key
 - Personalized Recommendations
 - Periodic Trends Packet Key User Reviews and Ratings

- Periodic Trends Packet Key and Bestseller Lists
- 5. Accessing Periodic Trends Packet Key Free and Paid eBooks
 - Periodic Trends Packet Key Public Domain eBooks
 - Periodic Trends Packet Key eBook Subscription Services
 - Periodic Trends Packet Key Budget-Friendly Options
- 6. Navigating Periodic Trends Packet Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends Packet Key Compatibility with Devices
 - Periodic Trends Packet Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Packet Key
 - Highlighting and Note-Taking Periodic Trends Packet Key
 - Interactive Elements Periodic Trends Packet Key
- 8. Staying Engaged with Periodic Trends Packet Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Packet Key
- 9. Balancing eBooks and Physical Books Periodic Trends Packet Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Packet Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends Packet Key
 - Setting Reading Goals Periodic Trends Packet Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends Packet Key
 - Fact-Checking eBook Content of Periodic Trends Packet Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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