Percent Composition and Molecular Formula Worksheet Key

- What's the empirical formula of a molecule containing 65.5% carbon, 5.5% hydrogen, and 29.0% oxygen? C₃H₃O mass = 55 g/mole
- If the molar mass of the compound in problem 1 is 110 grams/mole, what's the molecular formula?
 C₆H₆O₂
- What's the empirical formula of a molecule containing 18.7% lithium, 16.3% carbon, and 65.0% oxygen? Li₂CO₃
- If the molar mass of the compound in problem 3 is 73.8 grams/mole, what's the molecular formula?
 Li₂CO₃

Write the molecular formulas of the following compounds:

- A compound with an empirical formula of C₂OH₄ and a molar mass of 88 grams per mole. C₄O₂H₈
- 6. A compound with an empirical formula of C4H4O and a molar mass of 136 grams per mole. C8H8O2
- A compound with an empirical formula of CFBrO and a molar mass of 254.7 grams per mole. C₂F₂Br₂O₂
- 8. A compound with an empirical formula of C2H8N and a molar mass of 46 grams per mole. C2H8N

Answer the following questions:

- The percentage composition of acetic acid is found to be 39.9% C, 6.7% H, and 53.4% O. Determine the empirical formula of acetic acid. CH₂O
- 10. The molar mass for question #9 was determined by experiment to be 60.0 g/mol. What is the molecular formula? C₂H₄O₂
- 11. Aniline, a starting material for urethane plastic foams, consists of C, H, and N. Combustion of such compounds yields CO₂, H₂O, and N₂ as products. If the combustion of 9.71 g of aniline yields 6.63 g H₂O and 1.46 g N₂, what is its empirical formula? C₆H₇N
- 12. The molar mass of aniline is 93 g/mol. What is its molecular formula? CoH1N
- Calculate the mass percent of carbon, nitrogen and oxygen in acetamide, C₂H₅NO. %C 40.668 %H 8.533 %N 23.713 %O 27.086
- 14. A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms were produced. What is the empirical formula of the compound? PCIs
- When 2.5000 g of an oxide of mercury, (Hg_xO_y) is decomposed into the elements by heating, 2.405 g of mercury are produced. Calculate the empirical formula. Hg₂O
- 16. The compound benzamide has the following percent composition. What is the empirical formula?
 C = 69.40 % H= 5.825 % O = 13.21 % N= 11.57 % C₇H₇NO
- 17. A component of protein called serine has an approximate molar mass of 100 g/mole. If the percent composition is as follows, what is the empirical and molecular formula of serine?

C₃H₂NO₃ empirical formula C₄H₂NO₃ molecular formula

Percent Composition And Molecular Formula Answers

RJ Alexander

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