

ISOTOPES AND AVERAGE ATOMIC MASS

Name _____

Elements come in a variety of isotopes, meaning they are made up of atoms with the same atomic number but different atomic masses. These atoms differ in the number of neutrons.

The average atomic mass is the weighted average of all the isotopes of an element.

Example: A sample of cesium is 75% ^{133}Cs , 20% ^{132}Cs and 5% ^{134}Cs . What is its average atomic mass?

Answer: $.75 \times 133 = 99.75$

$.20 \times 132 = 26.4$

$.05 \times 134 = \underline{6.7}$

Total = 132.85 amu = average atomic mass

Determine the average atomic mass of the following mixtures of isotopes.

1. 80% ^{127}I , 17% ^{126}I , 3% ^{128}I

2. 50% ^{197}Au , 50% ^{198}Au

3. 15% ^{56}Fe , 85% ^{54}Fe

4. 99% ^1H , 0.8% ^2H , 0.2% ^3H

5. 95% ^{14}N , 3% ^{15}N , 2% ^{16}N

6. 98% ^{12}C , 2% ^{14}C

Periodic Table Chemistry If8766

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Table of Contents Periodic Table Chemistry If8766

1. Understanding the eBook Periodic Table Chemistry If8766
 - The Rise of Digital Reading Periodic Table Chemistry If8766
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Table Chemistry If8766
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Table Chemistry If8766
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Table Chemistry If8766
 - Personalized Recommendations
 - Periodic Table Chemistry If8766 User Reviews and Ratings
 - Periodic Table Chemistry If8766 and Bestseller Lists
5. Accessing Periodic Table Chemistry If8766 Free and Paid eBooks
 - Periodic Table Chemistry If8766 Public Domain eBooks
 - Periodic Table Chemistry If8766 eBook Subscription Services
 - Periodic Table Chemistry If8766 Budget-Friendly Options

6. Navigating Periodic Table Chemistry If8766 eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Table Chemistry If8766 Compatibility with Devices
 - Periodic Table Chemistry If8766 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Table Chemistry If8766
 - Highlighting and Note-Taking Periodic Table Chemistry If8766
 - Interactive Elements Periodic Table Chemistry If8766
8. Staying Engaged with Periodic Table Chemistry If8766
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Table Chemistry If8766
9. Balancing eBooks and Physical Books Periodic Table Chemistry If8766
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Table Chemistry If8766
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Periodic Table Chemistry If8766
 - Setting Reading Goals Periodic Table Chemistry If8766
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Table Chemistry If8766
 - Fact-Checking eBook Content of Periodic Table Chemistry If8766
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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