

- Key.
8. How many moles of copper atoms are needed to supply the same number of atoms as does 235 grams of iron?
- $$235 \text{ g} \left(\frac{1 \text{ mol}}{55.8 \text{ g}} \right) = 4.21 \text{ moles of Fe atoms}$$
- 4.21 mols of Cu atoms would have the same number of atoms.
9. How many grams of carbon are needed to supply twice as many atoms as does 27 grams of beryllium?
- $$27 \text{ g} \left(\frac{1 \text{ mol}}{9.0 \text{ g}} \right) = 3.0 \text{ mols Be atoms}$$
- twice atoms $2(3.0 \text{ mols}) = 6.0 \text{ mols}$
- $$6.0 \text{ mols C} \left(\frac{12.0 \text{ g}}{1 \text{ mol}} \right) = 72 \text{ g of carbon.}$$
10. 62.0 grams of rust is composed of 43.4 grams of iron. Determine the empirical formula for rust.
- $$43.4 \text{ g Fe} \rightarrow 43.4 \text{ g} \div 55.8 \text{ g/mol} = 0.78 \text{ mol Fe}$$
- $$62.0 - 43.4 \text{ g O} \rightarrow 18.6 \text{ g} \div 16.0 \text{ g/mol} = 1.16 \text{ mol O}$$
- mol ratio Fe : O
2 : 3
- Fe_2O_3
11. A solution of NaCl contains 35 g of solute dissolved in 100g of water. What is the molarity of the solution?
- $$\text{FW NaCl} = 58.5 \text{ g/mol}$$
- $$35 \text{ g} \div 58.5 \text{ g/mol} = 0.60 \text{ mol}$$
- $$100 \text{ g H}_2\text{O} \rightarrow 100 \text{ mL H}_2\text{O} \rightarrow 0.100 \text{ L}$$
- $$M = \frac{n}{V} = \frac{0.60 \text{ mol}}{0.100 \text{ L}} = 6.0 \text{ M}$$
12. The following three containers have identical capacities. Each contains Avogadro's number of molecules of a gaseous element. Fill in the blanks with the missing information.



Nelson Chemistry 11 Answer Key Chapter 8

Zhe Xu



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Properties of Polymers D.W. van Krevelen, 2012-12-02 Properties of Polymers Their Correlation with Chemical Structure Their Numerical Estimation and Prediction from Additive Group Contributions summarizes the latest developments regarding polymers their properties in relation to chemical structure and methods for estimating and predicting numerical properties from chemical structure In particular it examines polymer electrical properties magnetic properties and mechanical properties as well as their crystallization and environmental behavior and failure The rheological properties of polymer melts and polymer solutions are also considered Organized into seven parts encompassing 27 chapters this book begins with an overview of polymer science and engineering including the typology of polymers and their properties It then turns to a discussion of thermophysical properties from transition temperatures to volumetric and calorimetric properties along with the cohesive aspects and conformation statistics It also introduces the reader to the behavior of polymers in electromagnetic and mechanical fields of force The book covers the quantities that influence the transport of heat momentum and matter particularly heat conductivity viscosity and diffusivity properties that control the chemical stability and breakdown of polymers and polymer properties as an integral concept with emphasis on processing and product properties Readers will find tables that give valuable numerical data on polymers and include a survey of the group contributions increments of almost every additive function considered This book is a valuable resource for anyone working on practical problems in the field of polymers including organic chemists chemical engineers polymer processors polymer technologists and both graduate and PhD students

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