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formulae and equations e g in one mole of carbon c the number of atoms is the same as the number of molecules in one mole of carbon dioxide co 2 linking the mole and the atomic mass

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web for practical purposes a rounded version of the constant can be used in exams 6 02 x 10 23 mol 1 the reason we have mol 1 as the unit is because this is the number of entities per mole of substance for example one mole of sodium na contains $6 02 \times 10 23$ atoms of sodium one mole of hydrogen h 2 contains $6 02 \times 10 23$ molecules of hydrogen