

Chemistry 11

Chapter 4 – The MOLE

Atomic Masses and Avogadro's Hypothesis

Experimental work by the English chemist **John Dalton** (1766-1844) was concerned with how much of one element could combine with a given amount of another element. He suggested the following hypotheses:

- Molecules are made up of "atoms" of various elements
- If compound B contains twice the mass of element X as does compound A, then compound B must contain twice as many atoms as X
- Simple compounds are made up of only one atom of the two elements making up the compound (Dalton was aware that this hypothesis might be quite wrong but he had no way of knowing how many atoms were actually contained in a compound)

Dalton did not attempt to figure out the mass of an individual atom of any element. Instead he simply assigned an **arbitrary mass** to each element, assuming that hydrogen was the lightest element and therefore could be assigned a mass of "1". Since his experiments found that C was 6 times heavier than H, C was assigned a mass of 6. Similarly, O was 16 times heavier than H and was assigned a mass of "16". In this way, Dalton was able to calculate "relative masses" for several different elements.

Example: The reaction between 2.74 g of hydrogen gas and 97.26 g of chlorine gas makes 100 g of hydrogen chloride. Assuming hydrogen chloride contains one atom each of hydrogen and chlorine, the chlorine atom is

Exercise:

1. Assume a chemist working in 1810 (Mr. Pong – you're *THAT* old?) tries to duplicate some of Dalton's results, published a year or two earlier, and obtains the following results (expressed in modern mass unit, grams)

11.1 g of hydrogen gas react with 88.9 g of oxygen gas
46.7 g of nitrogen gas react with 53.3 g of oxygen gas
42.9 g of carbon react with 57.1 g of oxygen gas

Assuming a mass of "1" for hydrogen, calculate the relative masses of oxygen, nitrogen and carbon. (Don't be surprised if the values you calculate are not what you would expect from the periodic table. Not all the molecules produced in reactions involve 1:1 combinations of atoms, which was a problem plaguing early chemists for many years.

Dalton's atomic mass scale was partly in error because not all of the molecules he studied actually contained only one atom of each element. For example, it was now known that a molecule of hydrogen contains two atoms. Dalton's mass scale was just being introduced when the French chemist **Joseph Gay-Lussac** began to study how gases reacted.

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Principles of Applied Remote Sensing Siamak Khorram, Cynthia F. van der Wiele, Frank H. Koch, Stacy A. C.

Nelson, Matthew D. Potts, 2016-01-04 This textbook is one of the first to explain the fundamentals and applications of remote sensing at both undergraduate and graduate levels Topics include definitions and a brief history of payloads and platforms data acquisition and specifications image processing techniques data integration and spatial modeling and a range of applications covering terrestrial atmospheric oceanographic and planetary disciplines The policy and law issues of remote sensing and the future trends on the horizon are also covered Remote sensing is an exciting dynamic technology that is transforming the Earth sciences terrestrial atmospheric and marine as well as the practices of agriculture disaster response engineering natural resources providing evidence in legal cases and documented humanitarian crises and many other fields Increasingly understanding of these techniques will be central to a number of disciplines particularly as the technology advances The British National Bibliography ,1965 **The British National Bibliography** Arthur James Wells,1970

The United States Catalog Mary Burnham, Carol Hurd, 1928 *The United States Catalog* ,1928 **Nelson Chemistry, Atlantic Edition. Teacher's Resource (chapter 1-9)** Frank Jenkins, 1996 *Nelson Chemistry 11* Tompkins, Dick, Di Giuseppe, Maurice, 2002 Nelson Chemistry Frank Jenkins, 1996

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