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# THE NATURE OF GASES

# Section Review

#### Objectives

- Describe the assumptions of the kinetic theory as it applies to gases
- Interpret gas pressure in terms of kinetic theory.
- Define the relationship between Kelvin temperature and average kinetic energy

#### Vocabulary

- kinetic energy
- kinetic theory:
- gas pressure.

- Vacuum
- atmospherie pressure
- barometer

- pascal (Pa):
- · standard atmosphere (atm)

#### Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

|     |                           |                               |             | be particles of gas are relatively         |
|-----|---------------------------|-------------------------------|-------------|--|
|     |                           |                               |             | between particles are perfectly            |
|     |                           |                               |             | the simultaneous collisions of billions of |
|     | particles(with an object. | Barometers are used to measur | e press@re. | Standard conditions are defined as a       |
|     | temperature of            | and a pressure of             |             |  |
| ı.  | TG S                      |                               | .8          |  |
| 2.  |                           |                               | 9           |  |
| 3.  | 10                        | <u> </u>                      |             |  |
| 4.  |                           |                               |             |  |
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| 100 |                           |                               |             |  |
| 0.  |                           |                               |             |  |

- \_\_\_\_ 11. Atmospheric pressure is 760 mm Hg.
- \_\_\_\_\_ 12. The SI unit of pressure is the pascal.

# **Review Gases Section 1 Answers**

**Arshad Iqbal** 

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