

Honors Chem - Stoichiometry Quiz Summary Review

Stoichiometry

- the study of quantitative relationships in a chemical equation
- used to predict how much product will be formed
- used to determine the amount of reactants needed

EXAMPLE



The coefficients (underlined) of a balanced equation indicate the mole ratio of substances in a reaction.

How many moles of CO_2 will be produced when 8 moles of C_3H_8 combusts completely?

Step 1: Plan

We have to go from moles $\text{CO}_2 \rightarrow$ moles C_3H_8

Step 2: Calculate

$$\begin{aligned} 8 \text{ mol C}_3\text{H}_8 &\times 3 \text{ mol CO}_2 / 1 \text{ mol C}_3\text{H}_8 \\ &= 24 \text{ mol CO}_2 \end{aligned}$$

Limiting Reactant

- completely used by the reaction
- no remaining reactant at the end of the reaction
- determines the maximum amount of product formed
- balloon experiment flask D: vinegar (not enough to make the balloon any bigger)

Excess Reactant

- reactant left over at the end of the reaction
- does not completely react
- balloon experiment flask D: baking soda (was left over in the flask)

EXAMPLE

13.8 g of baking soda (NaHCO_3) was added to 0.0833 moles of vinegar ($\text{HC}_2\text{H}_3\text{O}_2$) according to the balanced equation below:



What is the maximum volume of CO_2 (g) that could be produced at STP?

(Useful GMFS (g/mol): NaHCO_3 - 84.0; $\text{HC}_2\text{H}_3\text{O}_2$ - 60.0)

Step 1: Plan

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University of California, Berkeley

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the University of California, Berkeley text.

Review Chemistry Honors Unit Stoichiometry:

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Stoichiometry Unit Project Luann Marie Decker, 1998

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