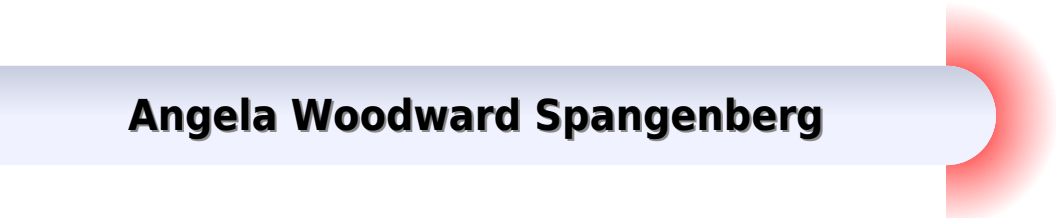


CHAPTER 6 REVIEW*Chemical Bonding***SECTION 4****SHORT ANSWER** Answer the following questions in the space provided.

1. _____ In metals, the valence electrons are considered to be
 - (a) attached to particular positive ions.
 - (b) shared by all surrounding atoms.
 - (c) immobile.
 - (d) involved in covalent bonds.
2. _____ The fact that metals are malleable and ionic crystals are brittle is best explained in terms of their
 - (a) chemical bonds.
 - (b) London forces.
 - (c) enthalpies of vaporization.
 - (d) polarity.
3. _____ As light strikes the surface of a metal, the electrons in the electron sea
 - (a) allow the light to pass through.
 - (b) become attached to particular positive ions.
 - (c) fall to lower energy levels.
 - (d) absorb and re-emit the light.
4. _____ Mobile electrons in the metallic bond are responsible for
 - (a) luster.
 - (b) thermal conductivity.
 - (c) electrical conductivity.
 - (d) All of the above.
5. _____ In general, the strength of the metallic bond _____ moving from left to right on any row of the periodic table.
 - (a) increases
 - (b) decreases
 - (c) remains the same
 - (d) varies
6. _____ When a metal is drawn into a wire, the metallic bonds
 - (a) break easily.
 - (b) break with difficulty.
 - (c) do not break.
 - (d) become ionic bonds.
7. Use the concept of electron configurations to explain why the number of valence electrons in metals tends to be less than the number in most nonmetals.

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Angela Woodward Spangenberg



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Table of Contents Review Chemical Bonding Section 1 Answers

1. Understanding the eBook Review Chemical Bonding Section 1 Answers
 - The Rise of Digital Reading Review Chemical Bonding Section 1 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Chemical Bonding Section 1 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Review Chemical Bonding Section 1 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review Chemical Bonding Section 1 Answers
 - Personalized Recommendations
 - Review Chemical Bonding Section 1 Answers User Reviews and Ratings
 - Review Chemical Bonding Section 1 Answers and Bestseller Lists
5. Accessing Review Chemical Bonding Section 1 Answers Free and Paid eBooks
 - Review Chemical Bonding Section 1 Answers Public Domain eBooks
 - Review Chemical Bonding Section 1 Answers eBook Subscription Services
 - Review Chemical Bonding Section 1 Answers Budget-Friendly Options

6. Navigating Review Chemical Bonding Section 1 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Review Chemical Bonding Section 1 Answers Compatibility with Devices
 - Review Chemical Bonding Section 1 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Review Chemical Bonding Section 1 Answers
 - Highlighting and Note-Taking Review Chemical Bonding Section 1 Answers
 - Interactive Elements Review Chemical Bonding Section 1 Answers
8. Staying Engaged with Review Chemical Bonding Section 1 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Chemical Bonding Section 1 Answers
9. Balancing eBooks and Physical Books Review Chemical Bonding Section 1 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Chemical Bonding Section 1 Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Review Chemical Bonding Section 1 Answers
 - Setting Reading Goals Review Chemical Bonding Section 1 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Review Chemical Bonding Section 1 Answers
 - Fact-Checking eBook Content of Review Chemical Bonding Section 1 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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