

Predicting Bond Types

TABLE 4.15 Predicting Bond Type from Electronegativity Differences

Molecule		Electronegativity Difference	Bond Type	Reason
H ₂	H—H	2.1 − 2.1 = 0	Nonpolar covalent	Less than 0.4
Cl ₂	Cl—Cl	3.0 − 3.0 = 0	Nonpolar covalent	Less than 0.4
HBr	$\overset{\delta^+}{\text{H}}-\overset{\delta^-}{\text{Br}}$	2.8 − 2.1 = 0.7	Polar covalent	Greater than 0.4, but less than 1.8
HCl	$\overset{\delta^+}{\text{H}}-\overset{\delta^-}{\text{Cl}}$	3.0 − 2.1 = 0.9	Polar covalent	Greater than 0.4, but less than 1.8
NaCl	Na ⁺ Cl [−]	3.0 − 0.9 = 2.1	Ionic	Greater than 1.8
MgO	Mg ²⁺ O ^{2−}	3.5 − 1.2 = 2.3	Ionic	Greater than 1.8

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