

AP® CHEMISTRY 2005 EXAMINATIONS CHEMISTRY SECTION

Question 10
(8 points)

As sample of $\text{C}_2\text{H}_2\text{Cl}_2$ is subjected to electrolysis, it is found that 0.010 moles of chlorine gas are formed for every 0.0050 moles of the sample. An electrochemical cell is constructed as follows:



Use the standard E° in brackets for the reduction half-reaction. The temperature of the cell is 25°C. Assume that the concentration of the electrolyte is 1.00 M.

Answer the following:

- (a) Write the half-reaction occurring at the cathode in the electrochemical cell. Indicate the number of electrons transferred in the half-reaction.
- (b) Write the half-reaction occurring at the anode in the electrochemical cell. Indicate the number of electrons transferred in the half-reaction.

1 point is earned for each half-reaction.

The cell is used to determine the standard reduction potential of the half-reaction:



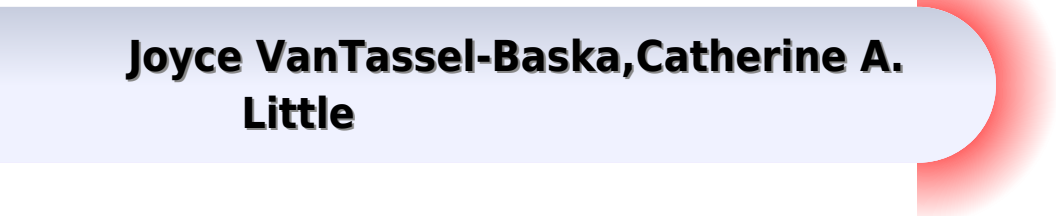
The cell is used to determine the standard reduction potential of the half-reaction. Write the half-reaction occurring at the anode in the electrochemical cell. Indicate the number of electrons transferred in the half-reaction.

- (a) Write the half-reaction occurring at the anode in the electrochemical cell. Indicate the number of electrons transferred in the half-reaction.
- (b) Write the half-reaction occurring at the cathode in the electrochemical cell. Indicate the number of electrons transferred in the half-reaction.

1 point is earned for each half-reaction.

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**Joyce VanTassel-Baska, Catherine A.
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