

## Review Sheet – Chemistry, Level 3 – Ch. 14: Liquids & Solids

Name: \_\_\_\_\_ ANSWERS \_\_\_\_\_

Date: \_\_\_\_\_

Period: \_\_\_\_\_

*This test covers materials in the textbook in Chapters 14 (pp. 456–491). You are always responsible for the topics on the previous tests.*

### Test Expectations

*The student should be able:*

- *Describe how the Kinetic-Molecular Theory of Matter predicts the properties of substances*
- *Analyze the forces and relative strengths of the forces between particles in a solid, liquid or gas and compare the substances in terms of basic properties such as melting point, boiling point, viscosity, surface tension, or vapor pressure*
- *Describe the type and properties of one of the four solids*
- *Draw or use phase diagrams or heating curves to describe phase changes*

1. Define each of the following terms clearly and completely.

(a) condensed state of matter –

(b) phase –

(c) phase diagram –

(d) heating curve –

(e) intramolecular force –

(f) intermolecular force –

(g) metallic bond –

(h) hydrogen bond –

(i) dipole-dipole force –

(j) dispersion force –

(k) permanent dipole –

(l) temporary dipole –

(m) induced dipole –

(n) viscosity –

(o) surface tension –

(p) water –

(r) crystalline solid –

(s) amorphous solid

(t) adhesion –

(u) cohesion –

(v) covalent-network solid –

(w) phase change –

(x) vapor pressure –

(y) heat of fusion –

(z) heat of vaporization –

(aa) sublimation –

(bb) deposition –

(cc) vaporization –

(dd) condensation –

(ee) phase change –

(ff) freezing –

(gg) melting –

(hh) Kinetic-Molecular Theory of Matter –

# Review Sheet Chemistry A Study Of Matter

**Kerry L. Hull, Barbara Janson Cohen**



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