

CHAPTER 6 REVIEW*Chemical Bonding***SECTION 4**

SHORT ANSWER Answer the following questions in the space provided.

1. _____ In metals, the valence electrons are considered to be
 - (a) attached to particular positive ions.
 - (b) shared by all surrounding atoms.
 - (c) immobile.
 - (d) involved in covalent bonds.
2. _____ The fact that metals are malleable and ionic crystals are brittle is best explained in terms of their
 - (a) chemical bonds.
 - (b) London forces.
 - (c) enthalpies of vaporization.
 - (d) polarity.
3. _____ As light strikes the surface of a metal, the electrons in the electron sea
 - (a) allow the light to pass through.
 - (b) become attached to particular positive ions.
 - (c) fall to lower energy levels.
 - (d) absorb and re-emit the light.
4. _____ Mobile electrons in the metallic bond are responsible for
 - (a) luster.
 - (b) thermal conductivity.
 - (c) electrical conductivity.
 - (d) All of the above.
5. _____ In general, the strength of the metallic bond _____ moving from left to right on any row of the periodic table.
 - (a) increases
 - (b) decreases
 - (c) remains the same
 - (d) varies
6. _____ When a metal is drawn into a wire, the metallic bonds
 - (a) break easily.
 - (b) break with difficulty.
 - (c) do not break.
 - (d) become ionic bonds.
7. Use the concept of electron configurations to explain why the number of valence electrons in metals tends to be less than the number in most nonmetals.

Review Sheet Chemical Bonding Answers

Karin Nielsen-Saines



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