

- Given the nuclear equation:
 $^{14}_7\text{N} + X \rightarrow ^{16}_8\text{O} + ^1_1\text{H}$
 What is particle X ?
 A) an alpha particle B) a beta particle
 C) a deuteron D) a triton
- The nucleus of a radium-226 atom is unstable, which causes the nucleus to spontaneously
 A) absorb electrons B) absorb protons
 C) decay D) oxidize
- Which group of nuclear emissions is listed in order of increasing charge?
 A) alpha particle, beta particle, gamma radiation
 B) gamma radiation, alpha particle, beta particle
 C) positron, alpha particle, neutron
 D) neutron, positron, alpha particle
- Which list of nuclear emissions is arranged in order from the least penetrating power to the greatest penetrating power?
 A) alpha particle, beta particle, gamma ray
 B) alpha particle, gamma ray, beta particle
 C) gamma ray, beta particle, alpha particle
 D) beta particle, alpha particle, gamma ray
- Base your answer to the following question on Given the nuclear equation:
 $^1_1\text{H} + X \rightarrow ^6_3\text{Li} + ^4_2\text{He}$
 The particle represented by X is
 A) ^9_4Li B) ^9_4Be C) $^{10}_5\text{Be}$ D) $^{10}_6\text{C}$
- Positrons are spontaneously emitted from the nuclei of
 A) potassium-37 B) radium-226
 C) nitrogen-16 D) thorium-232
- What is the decay mode of ^{37}K ?
 A) β^- B) β^+ C) γ D) α
- Which equation represents positron decay?
 A) $^{87}_{37}\text{Rb} \rightarrow ^0_{-1}\text{e} + ^{87}_{38}\text{Sr}$
 B) $^{237}_{92}\text{U} \rightarrow ^{233}_{90}\text{Th} + ^4_2\text{He}$
 C) $^{27}_{13}\text{Al} + ^4_2\text{He} \rightarrow ^{30}_{15}\text{P} + ^1_0\text{n}$
 D) $^{11}_6\text{C} \rightarrow ^0_{+1}\text{e} + ^{11}_5\text{B}$
- As ^{14}C decays to ^{14}N , the number of protons in the nucleus
 A) decreases B) increases
 C) remains the same
- Atoms of I-131 spontaneously decay when the
 A) stable nuclei emit alpha particles
 B) stable nuclei emit beta particles
 C) unstable nuclei emit alpha particles
 D) unstable nuclei emit beta particles
- Which statement best describes gamma radiation?
 A) It has a mass of 1 and a charge of 1.
 B) It has a mass of 0 and a charge of -1.
 C) It has a mass of 0 and a charge of 0.
 D) It has a mass of 4 and a charge of +2.
- Which nuclear decay emission consists of energy, only?
 A) alpha particle B) beta particle
 C) gamma radiation D) positron
- Which kind of nuclear radiation has high energy and no mass?
 A) alpha B) beta
 C) gamma D) neutron
- The half-life of a radioactive isotope is 20.0 minutes. What is the total amount of a 1.00-gram sample of this isotope remaining after 1.00 hour?
 A) 0.500 g B) 0.333 g
 C) 0.250 g D) 0.125 g

Half-life	Amount
20 min	1.00 g
40 min	0.50 g
60 min	0.25 g
80 min	0.125 g
- A radioactive element has a half-life of 2 days. Which fraction represents the amount of an original sample of this element remaining after 6 days?
 A) $\frac{1}{8}$
 B) $\frac{1}{2}$
 C) $\frac{1}{3}$
 D) $\frac{1}{4}$

2 days	1/1
4 days	1/2
6 days	1/4

Review Nuclear Chemistry Section 22 2 Answers

Wolfgang Guggemos



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