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CHAPTER 8 REVIEW

Chemical Equations and Reactions

SHORT ANSWER Answer the following questions in the space provided.

According to the law of conservation of mass, the total mass of the reacting substances is

- (a) always more than the total mass of the products.
- (b) always less than the total mass of the products.
- (c) sometimes more and sometimes less than the total mass of the products.
- (d) always equal to the total mass of the products.

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

Dute:

each question.		
2. In the reaction described by the wo	rd equation	
sodium	+ water -> sodium h	ydroxide + hydrogen gas
 a. sodium hydroxide is a product. 	ъ.	sodium is a product.
c. water is a reactant.	d.	Both (a) and (c)
3. Which of the following indicates th	at a chemical equation	on is balanced?
a. The numbers of atoms of each of	element are the same	on both sides of the equation.
 b. All of the coefficients are the sa 	ime.	
c. The numbers of molecules on e	ach side of the equat	ion are equal.
 The sums of the coefficients on 	each side of the equi	ation are equal.
 From a complete and correctly write 	tten chemical equatio	en, you can obtain the
 a. chemical formulas of the reacta 		
 b. relative amounts of the reactant 		
 c. physical states of the reactants i 	and products.	
d. All of the above		
5. To balance a chemical equation, yo	eu must	
 a. adjust the subscripts. 	b.	adjust the number of products.
 c. adjust the number of reactants. 	d.	adjust the coefficients.
6. The formation of a solid product in		
a. (aq). b. (g).	c. (/).	d. (r).
8. In the expression 3CO ₂ , the number	rs 3 and 2 are, respec-	
 a. a subscript and a coefficient. 	D	a coefficient and a subscript.
c. two subscripts.	d.	two coefficients.
9. What is the total number of molecu	iles that is represente $NH_4NO_2 \rightarrow N$	d by the following chemical equation? (Add up the coefficients.)
a. 3 b. 4	c. 8	d. 16
10. Which equation below violates the		
a. 2H₂ + O₂ → 2H₂O	De	KCl + Br → KBr + Cb
c. 2Fe ₂ O ₂ + 3C → 4Fe + 3CO ₂	d.	$Na_2CO_1 + 2HCI \rightarrow 2NaCI + H_2O + CO_2$
		ed equation for the reaction of iron and oxygen?
a. $2Fe + O_2 \rightarrow Fe_2O_1$	ь.	$Fe + 3O_2 \rightarrow Fe_2O_1$
c. 4Fe + 3O ₂ → 2Fe ₂ O ₃	-4	$3Fe + 3O_1 \rightarrow 2Fe_2O_1$
12. Which coefficients correctly balance	se the formula equati-	
	$KClO_2(r) \longrightarrow KC$	
a. 1, 1, 1 b. 1, 1, 3	c. 2, 2, 3	d. 2, 1, 1
13. Which equation is not balanced?		
a. $2H_2 + O_2 \rightarrow 2H_2O$	b.	$4H_2 + 2O_2 \rightarrow 4H_2O$
c. H ₂ + H ₂ + O ₂ → H ₂ O + H ₂ O	a.	2H₂ + O₂ → H₂O
14. In which type of reaction do two or		
a. synthesis (AKA combination)	Ъ.	decomposition
c. combustion	eff.	single-replacement
15. Which of the following is true rega	rding single-replacer	
a. Only one product is formed.	b.	There is only one reagent.
c. It requires oxygen.	d.	One element replaces another in a compound.

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K Morrison

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