

An Overview Of Nuclear Chemistry

AP Chemistry Review

A non-electrical, non-gravitational force called the nuclear force holds a nucleus together. Some nuclei are more stable than others. When a nucleus is unstable, it can attempt to increase its stability by altering its number of neutrons and protons. This is the process of radioactive decay.

TYPES OF NUCLEAR DECAY:

Alpha Emission:



In alpha decay, the nucleus emits a particle that has the same constitution as a helium nucleus, with 2 protons and 2 neutrons.



When a nucleus undergoes alpha decay:

- Subtract 4 from the mass #.
- Subtract 2 from the atomic #.

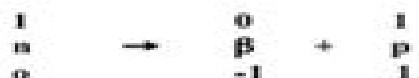


(remember, Thorium is identified by its atomic #, 90)

Beta Emission:



A beta particle is identical to an electron. In beta decay, the nucleus changes a neutron into a proton and an electron and emits the electron.



When a nucleus undergoes beta decay:

- The mass # remains the same.
- Add 1 to the atomic #.



Positron Emission:



A positron is like an electron with a positive charge. In positron emission, the nucleus changes a proton into a neutron and a positron, and emits the positron.



When a nucleus undergoes positron emission:

- The mass number remains the same.
- Subtract one from the atomic number.



Electron Capture:



In electron capture, the nucleus captures a low energy electron and combines it with a proton to form a neutron.

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