

Molecular Composition of Gases

SECTION: PART B (20)

QUESTION: ANSWER Answer the following questions in the space provided.

1. _____ The average speed of a gas molecule is most directly related to the _____.
 - (a) volume of the container
 - (b) pressure within the gas
 - (c) temperature of the gas
 - (d) number of molecules in the gas
2. _____ Which value below best represents the average molecular speed of a diatomic gas placed in the same temperature?
 - (a) The value associated with molecular mass is directly proportional.
 - (b) The value associated with molecular mass is directly proportional.
 - (c) The value associated with square root of molecular mass is directly proportional.
 - (d) The value associated with the square of molecular mass is directly proportional.
3. _____ Which combination of conditions corresponds to the following equation:

$$\text{volume} \propto \text{pressure} \propto \text{temperature}$$

 - a. _____ $\text{PV} = nRT$ is correct. The volume of H_2 is constant and the pressure of H_2 is constant.
 - b. _____ $\text{PV} = nRT$ is correct. The volume of H_2 is constant and the pressure of H_2 is constant.
4. _____ Which set of conditions is not a possible set of values for a gas?
 - a. _____ Mass, volume, molecular pressure, the weight of the gas is constant.
 - b. _____ Mass, volume, molecular pressure, the weight of the gas is constant.
 - c. _____ Molecular pressure, the weight, molecular mass is constant.

QUESTION: ANSWER Which gas has the greatest mass than the other two? Which has the least? Which will "float" above the other gases? Justify.

5. _____ Which gas has the greatest mass than the other two? Which has the least? Which will "float" above the other gases? Justify.

Review Molecular Composition Of Gases Short Answer

Sabine Zange



Review Molecular Composition Of Gases Short Answer:

Review Molecular Composition Of Gases Short Answer Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has been apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Review Molecular Composition Of Gases Short Answer**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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