



Physical Science

PDF Study Guide

Properties of Elements
 Study Guide: The Periodic Table, Ions, and...
 Questions and Answers
 1. Q: What are the main characteristics of metals?
 A: High electrical and thermal conductivity, and a shiny appearance.
 2. Q: Where are non-metals located on the periodic table?
 A: On the right side of the periodic table.
 3. Q: What is a noble gas?
 A: A gas that is chemically inert.
 4. Q: Name a characteristic of noble gases.
 A: They are inert and rarely form compounds.
 5. Q: What type of elements are found in the middle of the periodic table?
 A: Metals.
 6. Q: What type of elements are found on the far left of the periodic table?
 A: Alkali metals.
 7. Q: What type of elements are found on the far right of the periodic table?
 A: Noble gases.

Properties of Elements
 Study Guide: The Periodic Table, Ions, and Valence Electrons
 1. Overview of the Periodic Table
 The periodic table is a systematic arrangement of elements based on their atomic number, electron configurations, and recurring chemical properties. Elements are organized into horizontal rows (periods) and vertical columns (groups). The periodic table is divided into three main categories: metals, non-metals, and metalloids (metals in between).
 2. Metals, Non-Metals, and Metalloids
 Metals are located on the left side of the periodic table. They are characterized by their ability to conduct electricity and heat, their malleability, and their shiny appearance. Examples include iron, copper, and aluminum.
 Non-metals are located on the right side of the periodic table. They are generally poor conductors of electricity and heat, and they are brittle. Examples include oxygen, nitrogen, and carbon.
 Metalloids are located in between the metals and non-metals. They have properties that are intermediate between metals and non-metals. Examples include silicon, boron, and arsenic.
 3. Periodic Trends
 The periodic table shows several trends in the properties of elements. These trends include atomic size, ionization energy, and electronegativity. Atomic size increases from top to bottom and from left to right. Ionization energy decreases from top to bottom and from left to right. Electronegativity increases from bottom-left to top-right.

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 4. Valence Electrons
 Valence electrons are the electrons in the outermost shell of an atom. They are responsible for the chemical properties of an element. The number of valence electrons in an atom is equal to the number of electrons in the outermost shell. For example, sodium has one valence electron, while chlorine has seven.
 5. Ions
 Ions are atoms or molecules that have a net electrical charge. They are formed by the loss or gain of electrons. Cations are positively charged ions, and anions are negatively charged ions. For example, sodium loses an electron to form a sodium cation (Na⁺), and chlorine gains an electron to form a chloride anion (Cl⁻).

Properties of Elements
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 6. Periodic Trends (continued)
 The periodic table shows several trends in the properties of elements. These trends include atomic size, ionization energy, and electronegativity. Atomic size increases from top to bottom and from left to right. Ionization energy decreases from top to bottom and from left to right. Electronegativity increases from bottom-left to top-right.

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Educators will find a wealth of information and expert guidance to meet this need in *Resources for Teaching Elementary School Science* A completely revised edition of the best selling resource guide *Science for Children* *Resources for Teachers* this new book is an annotated guide to hands on inquiry centered curriculum materials and sources of help in teaching science from kindergarten through sixth grade Companion volumes for middle and high school are planned The guide annotates about 350 curriculum packages describing the activities involved and what students learn Each annotation lists recommended grade levels accompanying materials and kits or suggested equipment and ordering information These 400 entries were reviewed by both educators and scientists to ensure that they are accurate and current and offer students the opportunity to Ask questions and find their own answers Experiment productively Develop patience persistence and confidence in their own ability to solve real problems The entries in the curriculum section are grouped by scientific area Life Science Earth Science Physical Science and Multidisciplinary and Applied Science and by type core materials supplementary materials and science activity books Additionally a section of references for teachers provides annotated listings of books about science and teaching directories and guides to science trade books and magazines that will help teachers enhance their students science education *Resources for Teaching Elementary School Science* also lists by region and state about 600 science centers museums and zoos where teachers can take students for interactive science experiences Annotations highlight almost 300 facilities that make significant efforts to help teachers Another section describes more than 100 organizations from which teachers can obtain more resources And a section on publishers and suppliers give names and addresses of sources for materials The guide will be invaluable to teachers principals administrators teacher trainers science curriculum specialists and advocates of hands on science teaching and it will be of interest to parent teacher organizations and parents

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