

1. _____ **Ambident bases** _____, tend to release hydroxide ions (OH^-) in solution.
2. _____ **Bronsted-Lowry acids** _____ are substances that donate a proton.
3. _____ **Lewis acids** _____ are substances that can accept a pair of electrons.
4. The _____ **pOH** _____ scale describes the hydroxide ion concentration.
5. Acidic solutions have a high concentration of _____ **H_3O^+** _____ ions.
6. Solutions that have a pH above 7 are considered _____ **basic** _____.
7. A(n) _____ **amphoteric** _____ substance can act as both a Bronsted-Lowry acid and a Bronsted-Lowry base.
8. _____ **Buffers** _____ are solutions that resist pH changes.

Instructions: For each of the following questions, select the letter that best answers each question based on class notes and discussion. (1.5 points each)

- _____ 9. Which of the following is a property of acids?
- A. Have a bitter taste
 - B. Have to turn blue litmus paper red**
 - C. Have to turn red litmus paper blue
 - D. Have a slippery feeling
- _____ 10. What is neutralization reactions in aqueous solutions produce?
- A. A conjugate acid and conjugate base
 - B. A weak acid and a strong base
 - C. A salt and water**
 - D. A salt and a conjugate acid
- _____ 11. The scientist(s) who classified acids and bases by the ions they produce in solution were
- A. Sarrusent
 - B. Lewis
 - C. Bronsted and Lowry
 - D. Arrhenius**
- _____ 12. Using the Bronsted-Lowry model, strong acids

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