

Preliminary Chemistry: 8.2 The Chemical Earth

1. The living and non-living components of the Earth contain mixtures

- **Construct word and balanced formulae equations of chemical reactions as they are encountered**
 - **Identify the difference between elements, compounds and mixture in terms of particle theory**
1. Pure substances cannot be decomposed by simple physical separation techniques.
 2. Impure substances are mixtures. They can be separated into their components by various physical separation techniques.
 3. Elements are the simplest pure substances consisting of only one type of atom. They cannot be broken down or decomposed.
 4. Compounds are also pure substances. They are composed of two or more elements that are chemically bonded together. They are composed of a fixed number of atoms of each component element. They can be decomposed into their component elements.
 5. The properties of the compounds are quite different from their component elements.
 6. Atoms are the smallest particles of an element that can take part in a chemical reaction. Atoms of the same element or different elements group together to form more complex particles called molecules. Molecules are the smallest part of a pure substance that can exist separately. Noble gases are regarded as monatomic molecules.
 7. Homogeneous mixture is a mixture where all the particles are uniformly distributed.
 8. Heterogeneous mixture is a mixture where all the particles are not uniformly distributed.
- **Identify that the biosphere, lithosphere, hydrosphere and atmosphere contain examples of mixtures of elements and compounds**
1. Biosphere is an extensive region where living things are found. Lithosphere consists of the rigid, outer crust and the upper mantle. The hydrosphere is the watery part of the crust. It also extends into the atmosphere and lithosphere. Atmosphere is a layer of gas about 200 to 300 km thick that surrounds the planet.

	Element	Compound	Mixture
Biosphere	Oxygen, nitrogen	Carbohydrates, proteins, fats, vitamins	Cells
Atmosphere	Nitrogen, oxygen	Carbon dioxide, sulfur dioxide, carbon	Gases

		monoxide, nitrogen dioxide	
Hydrosphere	Oxygen, nitrogen	Carbon dioxide, magnesium chlorides, sodium chlorides	Sea water, ground water
Lithosphere	Gold, silver, platinum, carbon, copper	Oxides, sulfides, carbonates	Rocks, sand, soil, coal, oil

- **Identify and describe procedures that can be used to separate naturally occurring mixtures of:**
 - **solids of different sizes**
 - **solids and liquids**
 - **dissolved solids in liquids**
 - **liquids**
 - **gases**
- 1. Solids of different sizes
- Mined ore is first crushed and ground to produce finely powdered grains. The grinding separates the valuable minerals from each other and from the gangue. The crushed ore is then screened using large **wire-mesh sieves**, separating the valuable minerals from the gangue.
- Heavy mineral grains can be removed from the gangue by **sedimentation** and panning. In this technique, finely ground ore from the sieving process is mixed with water. Slurry is formed. This slurry is transported to **centrifuge** where the gravity causes the heavier particles to sediment while the lighter, finely ground material (that contains a greater proportion of the valuable minerals) is directed to the flotation tanks.
- In this technique, ground minerals are mixed with water, detergents and other oily chemicals to produce a froth to which the mineral grains adhere. The floating froth layer is scraped from the surface and removed from further treatment. This material is called the concentrate as it has a much higher percentage of valuable minerals than the original ore. The gangue does not float so sediments to the bottom of the vessel.
- Some iron mineral grains can be separated from the grains of crushed rock **by magnetic separation** technique. This technique can also be used to separate the magnetic ilmenite minerals from non-magnetic rutile minerals in sand mining operations. The mixture can pass

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