

The background of the cover art depicts a medieval battle scene. In the foreground, a knight in full plate armor, including a tall helmet with a visor, is mounted on a dark horse. He holds a sword aloft in his right hand and a large, orange-tinted shield in his left. Behind him, a line of soldiers on horseback follows. In the background, a stone castle is under attack, with smoke rising from the battlements and a large wooden siege engine, possibly a trebuchet, positioned to the right. The sky is filled with dark, swirling clouds.

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- \text{aq} + 2\text{Ag}^+ \text{aq}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver i ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 2ag aq 2f aq 2nh₄ aq cr₂o₇²⁻ aq ag₂cr₂o₇ s 2nh₄ aq

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