

Chapter 4 Practice Test – Answer Key

- _____ **d** 1) An element's properties can be predicted from its _____.
a) number of isotopes b) number of neutrons
c) atomic mass d) location in the periodic table
- _____ **c** 2) The _____ model of an atom is a ball of positive charge with negatively charged electrons embedded in it.
a) Dalton b) Rutherford c) Thomson d) Bohr
- _____ **a** 3) If a substance is _____, it can be pulled or drawn into a wire.
a) ductile b) malleable c) magnetic d) reactive
- _____ **b** 4) The spontaneous emission of radiation from an unstable atomic nucleus is called _____.
a) corrosion b) radioactivity c) nuclear fusion d) reactivity
- _____ **d** 5) A(n) _____ is a positively charged particle in an atom's nucleus.
a) electron b) neutron c) plasma d) proton
- _____ **a** 6) Which of the following is NOT a characteristic of most metals?
a) brittle b) ductile c) good conductor d) malleable
- _____ **b** 7) An element's _____ tells the number of protons in its nucleus.
a) atomic mass b) atomic number c) chemical symbol d) period
- _____ **c** 8) Dmitri Mendeleev created the first _____.
a) chemical reaction b) metal alloy c) periodic table d) semiconductor
- _____ **d** 9) The particles that are lost, gained, or shared in chemical reactions are _____.
a) plasmas b) protons c) neutrons d) electrons
- _____ **c** 10) A(n) _____ has some of the properties of both metals and nonmetals.
a) alloy b) alkaline earth metal c) metalloid d) noble gas
- _____ **b (diatomic molecule)** 11) A Tri-atomic Molecule consists of two nonmetal atoms bonded together.
a) True b) false
- _____ **b (alkali metals)** 12) The family of Alkaline Earth Metals is the most reactive group of metals in the periodic table.
a) True b) False
- _____ **a** 13) The Lanthanides and the Actinides are shown at the bottom of the periodic table in order to keep the table a reasonable size.
a) True b) False

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