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web google classroom for the following redox equation $\text{MnO}_4^{-} + \text{SO}_3^{2-} + 2\text{H}^{+} \rightarrow \text{Mn}^{2+} + \text{HSO}_4^{-}$ what will be the stoichiometric coefficient of HSO_4^{-} if the above equation is balanced in an acidic medium stuck use a hint report a problem do 4 problems

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^{-} + \text{Au} + \text{CN}^{-} + 2\text{e}^{-} + 2\text{H}_2\text{O}_2 \rightarrow 2\text{Au} + 2\text{HCN} + 2\text{H}_2\text{O}$ make electrons equal convert to basic solution $4\text{CN}^{-} + 2\text{Au} + 2\text{CN}^{-} + 2\text{e}^{-}$ multiplied by a factor of 2 $2\text{e}^{-} + 2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ the final answer

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation ni s ni a 2 a q 2 e now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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