7&8 • Atomic Structure & Periodicity

PRACTICE TEST

 $A = 2.18 \times 10^{-18} J$ $h = 6.626 \times 10^{-34} \text{ J/s}$ $R = 1.097 \times 10^{-7} \text{ m}^{-1}$ $c = 3.00 \times 10^8 \text{ m} \cdot \text{s}^{-1}$ mass of an electron = 9.11×10^{-31} kg

6.

- What wavelength corresponds to a frequency of 11. $8.22 \times 10^9 \text{ Hz?}$
 - a) 0.307 m

- d) 0.110 m
- b) 0.0365 m 500 (c) 27.4 m
- A radio station transmits at 110 MHz (110 x 106
 - Hz). What wavelength is this radio wave?
 - a) 3.65 x 10⁻⁵ m
- c) 3.81 x 10⁻⁵ m
- b) 3.30 m
- Which one of the following is NOT a proper unit for frequency?
 - a) Hz

b) s⁻¹

- Calculate the wavelength of the fourth line in the Balmer series (the visible series) of the hydrogen
 - a) 0.12334 m
 - d) 4.1029 x 10⁻⁷ m
 - b) 24.373 m
- e) 36.559 m
- c) 2.7353 x 10⁻⁷ m DE = 484+1519T
- What is the relationship between the energy of a photon of light and its frequency?
 - a) E = v

e) $E = \frac{V}{h}$

the hydrogen atom from the second energy level to the third energy level? 3 - E2 a) 1.52 x 10⁴ J d) 4.48 x 10⁻¹⁹ J

What is the energy needed to raise an electron in

- b) 3.63 x 10⁻¹⁹ J
- (e) 3.03 x 10⁻¹⁹ J
- c) 2.18 x 10-19 J -2.18 YOU W J
- What is the de Broglie wavelength of an electron moving at 80.0% the speed of light.

 - a) 3.03 x 10⁻¹² m c) 3.30 x 10¹¹ m
- λ: Δ = 2.42 x 10-12 m d) 1.59 x 10-25 m λ: Δ = 2.62 x 10-3 x 3.5 7.11 x 10 31 kg = .8 x 3.00 x 10 8
- What resultant is expected from the interference of the two waves shown below?



- b)
- Which quantum number determines the subshell occupied by an electron (s, p, d, f, etc.)?

C) ME,

d) esc.

Problem 23a Quantum Energy Answers

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