

Question	Expected Answers	Marks	Additional Guidance
3 a	measured pH > 1 OR $[H^+] < 0.1 \text{ (mol dm}^{-3}\text{)}$ ✓ $[H^+] = 10^{-0.94}$ ✓ $K_a = \frac{[H^+][CH_3CH_2COO^-]}{[CH_3CH_2COOH]}$ OR $\frac{[H^+]^2}{[CH_3CH_2COOH]}$ ✓ Calculate K_a from $\frac{[H^+]^2}{0.100}$ ✓	4	ALLOW C_2H_3 throughout question ALLOW $[H^+] < [CH_3CH_2COOH]$ OR $[H^+] < [HA]$ ALLOW measured pH is higher than expected ALLOW measured pH is not as acidic as expected ALLOW a quoted pH value or range > 1 and < 7 OR between 1 and 7 ALLOW $[H^+] = \text{antilog } -\text{pH}$ OR $[H^+] = \text{inverse log } -\text{pH}$ ALLOW $\frac{[H^+][A^-]}{[HA]}$ OR $\frac{[H^+]^2}{[HA]}$ IF K_a is NOT given and $K_a = \frac{[H^+]^2}{0.100}$ is shown, award mark for K_a also (I.e. $K_a = \frac{[H^+]^2}{0.100}$ is automatically awarded the last 2 marks)
b	Marks are for correctly calculated values. Working shows how values have been derived. $[H^+] = 10^{-13.46} = 3.47 \times 10^{-14} \text{ (mol dm}^{-3}\text{)}$ ✓ $[OH^-] = \frac{1.0 \times 10^{-14}}{3.47 \times 10^{-14}} = 0.29 \text{ (mol dm}^{-3}\text{)}$ ✓	2	ALLOW $3.467368505 \times 10^{-14}$ and correct rounding to 3.5×10^{-14} ALLOW 0.28840315 and correct rounding to 0.29, i.e. ALLOW 0.288 ALLOW alternative approach using pOH: pOH = $14 - 13.46 = 0.54$ ✓ $[OH^-] = 10^{-0.54} = 0.29 \text{ (mol dm}^{-3}\text{)}$ ✓ Correct answer gets BOTH marks

Ocr F325 June 2010 Question Paper

Terry C. Jones



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Table of Contents Ocr F325 June 2010 Question Paper

1. Understanding the eBook Ocr F325 June 2010 Question Paper
 - The Rise of Digital Reading Ocr F325 June 2010 Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr F325 June 2010 Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr F325 June 2010 Question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr F325 June 2010 Question Paper
 - Personalized Recommendations
 - Ocr F325 June 2010 Question Paper User Reviews and Ratings
 - Ocr F325 June 2010 Question Paper and Bestseller Lists
5. Accessing Ocr F325 June 2010 Question Paper Free and Paid eBooks
 - Ocr F325 June 2010 Question Paper Public Domain eBooks
 - Ocr F325 June 2010 Question Paper eBook Subscription Services
 - Ocr F325 June 2010 Question Paper Budget-Friendly Options

6. Navigating Ocr F325 June 2010 Question Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Ocr F325 June 2010 Question Paper Compatibility with Devices
 - Ocr F325 June 2010 Question Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr F325 June 2010 Question Paper
 - Highlighting and Note-Taking Ocr F325 June 2010 Question Paper
 - Interactive Elements Ocr F325 June 2010 Question Paper
8. Staying Engaged with Ocr F325 June 2010 Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr F325 June 2010 Question Paper
9. Balancing eBooks and Physical Books Ocr F325 June 2010 Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr F325 June 2010 Question Paper
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ocr F325 June 2010 Question Paper
 - Setting Reading Goals Ocr F325 June 2010 Question Paper
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr F325 June 2010 Question Paper
 - Fact-Checking eBook Content of Ocr F325 June 2010 Question Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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