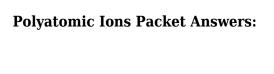
## List of Common Polyatomic Ions

Cations: Symbol	Name	Charge	Symbol	Name	Charge
NH <sub>4</sub> °	Ammonium	+1	$Hg_{2}^{-2}$	Dimercury (I)	+2
Anions:	5-4-00-0000	11.7 <u>12.4 1</u> 7.000 000 000 00			04444
Symbol	Name	Charge	Symbol	Name	Charge
$C_2H_3O_2$	Acetate	- 1	CO <sub>3</sub> ·2	Carbonate	-2
HCO <sub>3</sub>	Bicarbonate	- 1	CrO <sub>4</sub> -2	Chromate	-2
BrO <sub>3</sub>	Bromate	- 1	$Cr_2O_7^{-2}$	Dichromate	-2
BrO <sub>2</sub>	Bromite	- 1	$C_2O_4^{-2}$	Oxalate	-2
ClO <sub>4</sub>	Perchlorate	- 1	$O_2^{-2}$	Peroxide	-2
ClO <sub>3</sub>	Chlorate	- 1	SO <sub>4</sub> -2	Sulfate	-2
CIO <sub>2</sub>	Chlorite	- 1	SO <sub>3</sub> -2	Sulfite	-2
CIO	Hypochlorite	- 1	S <sub>2</sub> O <sub>3</sub> -2	Thiosulfate	-2
CN	Cyanide	- 1	SiO <sub>3</sub> 2	Silicate	-2
SCN	Thiocyanate	- 1	$MoO_4^{-2}$	Molybdate	-2
OH	Hydroxide	- 1	$WO_4^{-2}$	Tungstate	-2
NO <sub>3</sub>	Nitrate	- 1	BO <sub>3</sub> -3	Borate	-3
NO2	Nitrite	-1	$PO_4^{-3}$	Phosphate	-3
MnO <sub>4</sub>	Permanganate	- 1	$PO_3^{-3}$	Phosphite	-3
HSO.	Bisulfate	- 1	AsO <sub>4</sub> -3	Arsenate	-3
HSO <sub>3</sub>	Bisulfite	- 1	AsO <sub>3</sub> -3	Arsenite	-3
IO <sub>3</sub>	Iodate	- 1			
IO4	Periodate	- 1			
C <sub>6</sub> H <sub>5</sub> COO	Benzoate	-1			

Symbol	Name	Charge
$H_{2}PO_{4}^{-1}$	Dihydrogen phosphate	-1
$HPO_4^{-2}$	Hydrogen phosphate	-2
Fe(CN) <sub>6</sub> -3	Hexacyannoferrate (III)	-3
Fe(CN) <sub>6</sub> -4	Hexacyannoferrate (II)	-4

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**M** Mosston



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the boiling point of an organic liquid is the temperature at which its vapour pressure equals the atmospheric pressure over the liquid or it is the temperature at which the vapour and liquid phases are in equilibrium at agiven pressure the boiling point is considered as a criterion of purity of a compound and is useful for identification of

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given that the boiling points of ethanol and methanol 78 c and 64 c respectively are far off in a significant way this seems like a valid conclusion conclusion thermometer calibrations are necessary in order to collect reliable information as precise calculations are imperative to gathering correct results

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melting point 1 a push the open end of a capillary tube into the powdered acetamide b move the powder to the closed end of the capillary tube by tapping it on the table repeat until the the powdered acetamide occupies 1 2 mm of the capillary tube

end c

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discussion and conclusion the technique recrystallization was used because it is a simple method to purify a solid compound the percentage yield for benzoic acid and acetanilide acid being high indicated the initial sample had a low concentration of impurities with the melting point being in a specific

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determination of boiling points introduction the boiling point of a compound is the temperature at which it changes from a liquid to a gas this is a physical property often used to identify substances or to check the purity of the compound it **boiling point determination lab report introduction the** - Jul 19 2023

clean up solid and liquid spills immediately true g conclusions after conducting the boiling point determination lab it can be concluded that the unknown substance 193 is isopropyl alcohol on account of its observed boiling point of 88 c

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conclusion after completing this experiment the melting point of three different unknown compounds was obtained using the mel temp apparatus compound a had a melting point range of 114 6 115 5 c thus revealing it to be acetanilide compound b had a melting point range of 122 6 124 5 c thus

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melting point can be used as the identity and purity of an organic solid however it is not the same as the boiling point the boiling point is a characteristic unique of pure liquids the boiling point is the temperature at which the vapor pressure of a liquid is equal to the atmospheric pressure

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exploring boiling points abstract in this experiment we are given a set of straight chain alkanes and a set of compounds tocompare by following the trendline of the straight chain alkanes it could be concluded that asthe molar mass of the straight chain alkanes increases their boiling points increase as well

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50 c so in order to convert an experimental measurement taken in calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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