

## Physical Science Standards:

**SPS1:** Students will investigate our current understanding of the atom.

### a. Examine the structure of the atom in terms of

- proton, electron, and neutron locations.
  - Proton has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Protons are also the elements \_\_\_\_\_ number of the identity of the element, so protons are found by looking at the atomic number
  - Neutron has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Neutrons are found by subtracting the \_\_\_\_\_ Number from the \_\_\_\_\_ Mass or the BIG #- SMALL #.
  - Electron has a charge of \_\_\_\_\_ and is located in the \_\_\_\_\_ of the atom.
    - Electrons are equal to the number of \_\_\_\_\_ in a neutral atom.
  - The nucleus is made up of these 2 particles, \_\_\_\_\_ and \_\_\_\_\_.
  - The electrons are located in this region of the atom, known as the \_\_\_\_\_.
  - The electron cloud is divided in to energy levels or energy shells, the first level will hold \_\_\_\_\_ electrons and the second level will hold \_\_\_\_\_ electrons.
  - The valence shell or valence level is the \_\_\_\_\_ energy shell of the atom.
- atomic mass and atomic number.
  - Atomic number is equal to the number of \_\_\_\_\_ in the atom and is also the atoms identity—it's the small number of the periodic table.
  - Atomic mass is equal to the average mass of all the isotopes of a particular element—it the big number on the periodic table.
- atoms with different numbers of neutrons (isotopes).
  - Define Isotope \_\_\_\_\_.
  - Ex. Carbon-12 and Carbon-14
    - Carbon 12 has \_\_\_\_\_ P \_\_\_\_\_ N \_\_\_\_\_ E
    - Carbon 14 has \_\_\_\_\_ P \_\_\_\_\_ N \_\_\_\_\_ E
- explain the relationship of the proton number to the element's identity.
  - The \_\_\_\_\_ number is equal to the number of \_\_\_\_\_ in the atom which is how the element is identified.
  - Ex. Nitrogen has an atom number of \_\_\_\_\_, which means that is has \_\_\_\_\_ protons.

molybdenum	← element name
42	← atomic number number of protons (Z)
Mo	← atomic symbol
95.94	← atomic mass A (this is an average mass)



- Number of Protons = Atomic Number
- Number of Electrons = Number of Protons = Atomic Number
- Number of Neutrons = Mass Number - Atomic Number

# Physical Science Eoct Study Guide Forsyth

**L Darling-Hammond**



## **Physical Science Eoct Study Guide Forsyth:**

## **Physical Science Eoct Study Guide Forsyth** Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has be more apparent than ever. Its power to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Physical Science Eoct Study Guide Forsyth**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound affect our existence. Throughout this critique, we shall delve to the book is central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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