

Physical Science Final Exam REVIEW

NAME: _____ CLASS: _____

The following Physical Science competencies included in this exam packet:

1. Understand and apply concepts of the scientific method, introductory mathematical modeling, the use of scientific equipment, and scientific reporting in solving basic physics and chemistry problems.
2. Model, apply concepts, and solve problems related to the motion of objects and underlying forces.
3. Demonstrate, apply concepts, and solve problems related to the properties of light and sound waves.
4. Show the relationships between electricity and magnetism, apply those concepts and describe their roles in technology.
5. Make use of thermodynamic principles to show how matter and energy change within a system.
6. Show how the evolution of the atomic model explains the structure of the periodic table, the nature of atoms and make use of the periodic table to examine properties of elements.
7. Demonstrate their understanding of nuclear radiation, atomic energy, and the evolution of the universe.
8. Illustrate, apply concepts, and solve problems related to molecular bonding and chemical reactions.
9. Describe, apply concepts, and solve problems related to acids, bases, and carbon-based chemistry.

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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^- + \text{Au} \rightarrow 2\text{e}^- + 2\text{H}_2\text{O} + 2\text{H}_2\text{O}_2$ make electrons equal convert to basic solution $4\text{CN}^- + 2\text{Au} \rightarrow 2\text{e}^- + 2\text{H}_2\text{O} + 2\text{H}_2\text{O}_2$ multiplied by a factor of 2 $2\text{e}^- + 2\text{H}_2\text{O} + 2\text{H}_2\text{O}_2$ the final answer

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web aug 29 2023 the first step to balance any redox reaction is to separate the reaction into half reactions the substance being reduced will have electrons as reactants and the oxidized substance will have electrons as products usually all reactions are written as reduction reactions in half reaction tables

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web a redox equation can be balanced using the following stepwise procedure 1 divide the equation into two half reactions 2 balance each half reaction for mass and charge 3 equalize the number of electrons transferred in each half reaction 4

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web the half equation method separates the oxidation and reduction of a redox reaction in half reactions overall scheme for the half reaction method step 1 split reaction into half reactions reduction and oxidation step 2 balance the charge or oxidation number with electrons step 3 balance c e o by adding c e h₂o

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web when balancing equations for redox reactions occurring in basic solution it is often necessary to add oh⁻ ions or the oh⁻ h₂o pair to fully balance the equation in this video we ll walk through this process for the reaction between clo⁻ and cr oh₄⁻ in basic solution created by jay

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web bonus problem here is a balanced net ionic equation $3\text{Br}_2 + 6\text{OH}^- \rightarrow 5\text{Br}^- + 3\text{H}_2\text{O}$ the above equation takes place in two stages first bromine reacts with OH^- ions to form bromide ions and BrO^- ions second the BrO^- ions react to form bromide ions and bromate BrO_3^- ions

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web balancing redox reactions in acidic solution problems 1 10 problem 1 $\text{Cr}_2\text{O}_7^{2-} + \text{Fe}^{2+} + \text{Cr}^{3+} + \text{Fe}^{3+}$ solution 1 balanced half reactions $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$ 2 equalize the electrons $6\text{e}^- + 14\text{H}^+ + \text{Cr}_2\text{O}_7^{2-} \rightarrow 2\text{Cr}^{3+} + 7\text{H}_2\text{O}$ $6\text{Fe}^{2+} \rightarrow 6\text{Fe}^{3+} + 6\text{e}^-$ multiplied by a factor of 6

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web balance each half reaction in basic solution 4 5 6 7 $\text{Cr}_2\text{O}_7^{2-} + \text{NO}_2 + \text{SO}_4^{2-} + \text{MnO}_2 + \text{Cr}^{3+} + \text{NO}_3^- + \text{SO}_2 + \text{Mn}_2\text{O}_3$ balance each redox reaction in acid solution using the half reaction method 8 $\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ + \text{TeO}_3^{2-} + 2\text{H}^+ + \text{NO}_2 + \text{O}_2 + \text{Cr}^{3+} + \text{Te} + 10\text{H}^+ + \text{ReO}_4^- + 11\text{H}^+ + \text{PbO}_2 + \text{IO}_3^- + \text{Pb} + 12\text{H}^+ + \text{AsO}_3^{3-} + \text{Re} + \text{IO}_3^- + \text{H}_2\text{AsO}_4^-$

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equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding h₂ o₂ h and e in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively c₆h₈o₆ 2h₂no₂ c₆h₆o₆ 2h₂o 2no choose 1 answer c₆h₈o₆ and no₂ a c₆h₈o₆ and no₂ no 2 and c₆h₈o₆ b no₂ and c₆h₈o₆ h and no₂ c h and

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