

Question	Answer	Marks	Guidance
5 (a)	<p>F_2 forces mark F_2 has van der Waals' (forces) OR F_2 has induced dipole attractions OR interactions OR F_2 has temporary OR instantaneous dipole(-dipole) attraction OR interactions ✓</p> <p>HCl forces mark HCl has permanent dipole(-dipole) attractions OR interactions ✓</p> <p>Comparison of strength of forces between molecules mark intermolecular force in HCl is stronger than that in F_2 OR permanent dipoles are stronger (than induced dipoles) ✓</p> <p>Boiling point mark more energy is required to break stronger (intermolecular) forces ✓</p>	4	<p>Use annotations ie ticks crosses ECF ^ etc for this part</p> <p>ALLOW vdWs for van der Waals' IGNORE F_2 has covalent bond for this mark IGNORE F_2 has 'intermolecular forces'</p> <p><i>Quality of written communication: 'dipole(s)' spelled correctly and used in context for the second marking point</i> IGNORE HCl has 'intermolecular forces' IGNORE van der Waals' forces in HCl/ DO NOT ALLOW hydrogen bonding DO NOT ALLOW ionic bonding</p> <p>Look for strength of force comparison anywhere in the answer ALLOW ECF for hydrogen bonding in HCl being stronger than the stated intermolecular forces in F_2 BUT DO NOT ALLOW this mark if HCl or F_2 has covalent bonds broken OR if HCl has ionic bonds broken (the question asks for forces between molecules) IGNORE HCl has stronger van der Waals' (forces) than F_2 (as they both have the same number of electrons)</p> <p>DO NOT ALLOW fourth mark if covalent bonds are broken in HCl or F_2 OR if ionic bonds are broken in HCl</p> <p>IGNORE 'heat' but ALLOW 'heat energy'</p>

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Siddappa N.Byrareddy



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