

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance Basic SI Unit? kg
 - Volume of an object? ruler or water displacement method Basic SI Unit? cm³ or ml
 - Length of an object? Ruler Basic SI Unit? meter
 - Density of an object? graduated cylinder Basic SI Unit? ml
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 25 g and has a volume of 3 cm³? 8g/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has.
solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of neutrons and protons.

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Thomas Griffiths



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Table of Contents Place Science Exam Study Guide

1. Understanding the eBook Place Science Exam Study Guide
 - The Rise of Digital Reading Place Science Exam Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Place Science Exam Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Place Science Exam Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Place Science Exam Study Guide
 - Personalized Recommendations
 - Place Science Exam Study Guide User Reviews and Ratings
 - Place Science Exam Study Guide and Bestseller Lists
5. Accessing Place Science Exam Study Guide Free and Paid eBooks
 - Place Science Exam Study Guide Public Domain eBooks
 - Place Science Exam Study Guide eBook Subscription Services
 - Place Science Exam Study Guide Budget-Friendly Options

6. Navigating Place Science Exam Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Place Science Exam Study Guide Compatibility with Devices
 - Place Science Exam Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Place Science Exam Study Guide
 - Highlighting and Note-Taking Place Science Exam Study Guide
 - Interactive Elements Place Science Exam Study Guide
8. Staying Engaged with Place Science Exam Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Place Science Exam Study Guide
9. Balancing eBooks and Physical Books Place Science Exam Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Place Science Exam Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Place Science Exam Study Guide
 - Setting Reading Goals Place Science Exam Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Place Science Exam Study Guide
 - Fact-Checking eBook Content of Place Science Exam Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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