



NATIONAL  
SENIOR CERTIFICATE

GRADE 11

NOVEMBER 2020

PHYSICAL SCIENCES P2  
(CHEMISTRY)  
EXEMPLAR

MARKS: 100  
TIME: 1 hour



The question paper consists of 10 pages, including 4 extra sheets.  
Use blue or black ink for writing. Do not use correction fluid.

# QUESTION 3

## PHYSICAL SCIENCES PAPER 2

## INTERMOLECULAR FORCES

**GRADE 11**

**NOVEMBER 2020**

Examiner's use: \_\_\_\_\_ Physical Sciences P2

QUESTION 3 starts on a NEW page.

3.1 The boiling points in the table below were obtained during an investigation carried out to measure the boiling points of pure 1-hexanol. The table also lists the  $\Delta H_{\text{vap}}$  and  $\Delta H_{\text{fusion}}$  for 1-hexanol in kJ mol<sup>-1</sup> and 1 mol of 1-hexanol.

| Compound | Molecular mass (M <sub>r</sub> ) | Boiling point (°C) |
|----------|----------------------------------|--------------------|
| A        | 102                              | 16,2               |
| B        | 102                              | 21                 |

- 3.1.1 Define the term boiling point. (2)
- 3.1.2 Is pure 1-hexanol only composed of A and B or of A and B and one of the other compounds? (1)
- 3.1.3 Name the type of intermolecular force that exists between molecules of both compounds A and B and is the prime reason of their volatility. (1)
- 3.1.4 What factor of the compounds A and B is responsible for compound A having a higher boiling point? (1)
- 3.1.5 What compound (A or B) will have the higher vapour pressure? (1)
- 3.1.6 What compound (A or B) will have the higher surface tension? (1)

Compound B (M<sub>r</sub> 102) will show in the table, has a smaller molecular mass than both compounds A and B but has a relatively higher boiling point of 16,2 °C.

- 3.1.6 Explain why the boiling point of compound B is 16,2 °C (lower) than that of compound A and B by referring to the  $\Delta H_{\text{vap}}$  and intermolecular forces involved. (2)

3.2 Solid 1-hexanol (C<sub>6</sub>H<sub>14</sub>O) is added to a beaker of water. At room temperature (25 °C) and under 1 atmosphere the solid is shown in the diagram below.



- 3.2.1 Is solid 1-hexanol (C<sub>6</sub>H<sub>14</sub>O) or H<sub>2</sub>O, the major solvent? (1)
- 3.2.2 Explain the process in QUESTION 3.2.1 about the solubility in the intermolecular forces involved. (2)

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