

PHYSICAL SCIENCES – GRADE 10

CHAPTER – 1 Chemical Reactions and Equations

- q **Chemical Reaction** : – Whenever a chemical change occurs we can say that a chemical reaction has taken place
eg – Food gets digested in our body
 – Rusting of iron.
- q **Chemical Equation** :– A chemical reaction can be expressed symbolically by using chemical equation
eg magnesium is burnt into air to form magnesium oxide can be represented as
$$\text{Mg} + \text{O}_2 \rightarrow \text{MgO}$$
 - We can observe or recognise a chemical reaction by observing change in state, colour, by evolution of gas or by change in temperature.
- q Physical state of the reactant and products are mentioned to make chemical reaction more informative. eg we use (g) for gas, (l) for liquid, (s) for solid and (aq) for aqueous.
- q **Balancing Equation** :– We balance the chemical equation so that no. of atoms of each element involved in the reaction remain same at the reactant and product side.
eg $\text{Fe} + \text{H}_2\text{O} \rightarrow \text{Fe}_3\text{O}_4 + \text{H}_2$ can be written as
$$3\text{Fe(s)} + 4\text{H}_2\text{O(g)} \rightarrow \text{Fe}_3\text{O}_4\text{(s)} + 4\text{H}_2\text{(g)}$$
- q **Combination Reaction** :– The reaction in which two or more substances combine to form a new single substance

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Table of Contents Physical Science Grade 10 Page 1

1. Understanding the eBook Physical Science Grade 10 Page 1
 - The Rise of Digital Reading Physical Science Grade 10 Page 1
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Grade 10 Page 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Grade 10 Page 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Grade 10 Page 1
 - Personalized Recommendations
 - Physical Science Grade 10 Page 1 User Reviews and Ratings
 - Physical Science Grade 10 Page 1 and Bestseller Lists
5. Accessing Physical Science Grade 10 Page 1 Free and Paid eBooks
 - Physical Science Grade 10 Page 1 Public Domain eBooks
 - Physical Science Grade 10 Page 1 eBook Subscription Services
 - Physical Science Grade 10 Page 1 Budget-Friendly Options

6. Navigating Physical Science Grade 10 Page 1 eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science Grade 10 Page 1 Compatibility with Devices
 - Physical Science Grade 10 Page 1 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Grade 10 Page 1
 - Highlighting and Note-Taking Physical Science Grade 10 Page 1
 - Interactive Elements Physical Science Grade 10 Page 1
8. Staying Engaged with Physical Science Grade 10 Page 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Grade 10 Page 1
9. Balancing eBooks and Physical Books Physical Science Grade 10 Page 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Grade 10 Page 1
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Grade 10 Page 1
 - Setting Reading Goals Physical Science Grade 10 Page 1
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Grade 10 Page 1
 - Fact-Checking eBook Content of Physical Science Grade 10 Page 1
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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