

Name \_\_\_\_\_

# Physical Science Chemical Bonding Study Guide

## A. Covalent & Ionic Compounds/Bonds

1. Which ion accepts electrons, and which from it? (anion/cation) \_\_\_\_\_  
which cation it gains and electron it loses from it? \_\_\_\_\_ (anion/cation)
2. How placed in the halogens group has \_\_\_\_\_ valence electrons, and number is \_\_\_\_\_ 1 smaller in Group 16. This is because of their \_\_\_\_\_ rules.
3. How \_\_\_\_\_ bond in the addition between oppositely charged ions. (ionic covalent bond) is \_\_\_\_\_
4. On a molecular level, atoms \_\_\_\_\_ molecules, formed by any bond bond, atoms \_\_\_\_\_ molecules.
5. (Electron group) Which side from left shared completely by it completely bond is to \_\_\_\_\_, and others (group) shared equally to the \_\_\_\_\_  
How can predict if a compound is ionic, polar covalent, or nonpolar covalent by  
looking at the \_\_\_\_\_
6. How atoms can have up to \_\_\_\_\_ valence electrons. This usually happens in the periodic table that all have 8 valence electrons, or called \_\_\_\_\_
7. Atoms can gain, lose, or share \_\_\_\_\_ by forming ionic bonds.

## B. Formulas for Ionic Compounds – Put in the right order

Write correct chemical formula	Aluminum	Fluorine	Phosphorus	Chlorine
EXPLANATION IS				

## C. Write the chemical formulas for each of the following using subscripts

Think your own about the Lewis structures, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.

# Physical Science Chemical Bonding Study Guide

**Clemens Wendtner**



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