

Question	Answer	Marks	Guidance								
1 (a) (i)	<p>Atom(s) of an element</p> <p><b>AND</b></p> <p>with different numbers of neutrons (and with different masses) ✓</p>	1	<p><b>ALLOW</b> for 'atoms of an element':</p> <p>Atoms of the same element</p> <p>OR atoms with the same number of protons</p> <p>OR atoms with the same atomic number</p> <p><b>IGNORE</b> 'different relative atomic masses'</p> <p><b>IGNORE</b> different mass number</p> <p><b>IGNORE</b> same number of electrons</p> <p><b>DO NOT ALLOW</b> different numbers of electrons</p> <p><b>DO NOT ALLOW</b> 'atoms of elements' for 'atoms of an element'</p> <p><b>DO NOT ALLOW</b> 'an element with different numbers of neutrons' (ie atom(s) is essential)</p>								
	<p>(ii)</p> <table border="1" data-bbox="208 611 846 685"> <tr> <td>Protons</td><td>Neutrons</td><td>Electrons</td><td></td></tr> <tr> <td>74</td><td>110</td><td>74</td><td>✓</td></tr> </table>	Protons	Neutrons	Electrons		74	110	74	✓	1	
Protons	Neutrons	Electrons									
74	110	74	✓								
	(iii) $^{12}\text{C}$ OR C-12 OR carbon 12 OR carbon-12 ✓	1	<b>IGNORE</b> $1/12^{\text{th}}$ <b>AND</b> amu								
(b)	<p>(i) (Oxidised):</p> <p>H (oxidation number has increased) from H = 0 to H = +1 ✓</p> <p>(Reduced):</p> <p>W (oxidation number has decreased) from W = +6 to W = 0 ✓</p>	2	<p><b>ALLOW</b> 6+ OR 6 OR 1+ OR 1</p> <p><b>ALLOW</b> one mark for correct oxidation number changes</p> <p>H = 0 to H = +1 <b>AND</b> W = +6 to W = 0</p> <p><b>ALLOW</b> oxidation states written above the equation if not seen in the text <b>BUT IGNORE</b> oxidation states written above the equation if seen in the text</p> <p><b>ALLOW</b> for one mark: (Oxidised) H has increased by 1 <b>AND</b> (Reduced) W has decreased by 6</p> <p><b>IGNORE</b> <math>\text{WO}_3</math> is reduced</p> <p><b>IGNORE</b> references to electron loss / gain if correct</p> <p><b>DO NOT ALLOW</b> incorrect references to electron loss / gain</p> <p><b>DO NOT ALLOW</b> 'H oxidised and W reduced' without reference to oxidation number changes</p>								

# Ocr F321 Jan 2013 Question Paper

**JG Myers**



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