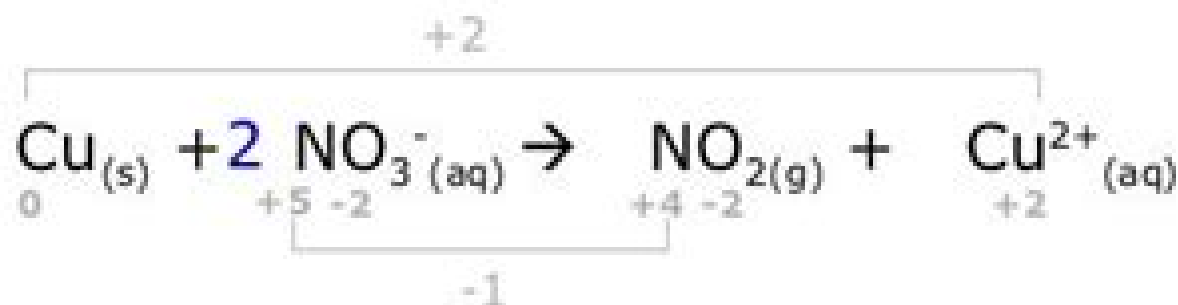


OXIDATION NUMBER METHOD

Example #3: In **acidic** solution

1. Assign oxidation numbers to each element.
2. Identify the increase and decrease in oxidation numbers.
3. Determine the smallest whole-number ratio of the oxidized and reduced elements so that the total increase in oxidation numbers equals the total decrease in oxidation numbers.
4. Use the determined ratio from step 3 to balance the oxidized and reduced reactants.
5. Balance the other elements/compounds by inspection.
6. Add H_2O , H^+ , and OH^- as needed to balance reactions in acidic or basic solution (recall the half-reaction method rules)



Ratio = 1:2

Each Cu atom gains 2, $\therefore$ 1 Cu gains 2

Each N atom loses 1, $\therefore$ 2 N loses 2

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