

STUDY GUIDE 11: Oxidation and Reduction (Chapter 12)

Remember: Oxidation = Loss of Electrons

Oxidation: loss of electrons; decrease in O.N. number (including 0, 0, 0 bonds)

Reduction: gain of electrons; increase in O.N. number; decrease in O.N. number



Be careful: some compounds contain both oxidizing and reducing agents, but in single reactions



REDOX REACTIONS Redox reactions involve the

Transfer of

Electrons. For any redox reaction, the sum of the oxidation numbers of the reactants must equal the sum of the oxidation numbers of the products.



Remember: Redox reactions are important in many areas of chemistry, including biology, medicine, and industry.



Oxidation And Reduction Study Guide

CL Gary



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