

## GROUP 2 ELEMENTS AND THEIR COMPOUNDS

- show their outermost electron placed on a subshell
- readily react to lose 2 electrons
  - $\text{M} \rightarrow \text{M}^{2+} + 2\text{e}^-$

### Trends down the group:

- 1<sup>st</sup> ionisation energy decreases
  - Distance between outermost electron and nucleus increases
  - More shielding
  - Counteractive effect of increasing nuclear charge
- Atomic radius increases
- Melting point generally decreases
  - Much lower than all transition metals

### Physical properties:

- Good conductors of heat and electricity
  - Sea of delocalised electrons
- Form white compounds
  - Transition metals form coloured compounds
- Oxidation number of +2
- All compounds are toxic or almost toxic
  - Be compounds show some covalent character because small highly charged cation polarises anion shell of anion
- Oxides / hydroxides are basic
  - React with acids (proton acceptors)
  - Except:
    - $\text{BeO}$
    - $\text{Be(OH)}_2$  which is amphoteric
- Density =  $2.3\text{--}3.7\text{ g/cm}^3$ 
  - Average density of transition metals is about  $7\text{ g/cm}^3$
- Compared to group 2 metals:
  - Harder + denser
  - Higher melting point
    - Metallic bond is stronger (higher charge on cation & more electrons delocalised)

### Chemical properties:

- Strong reducing agents
- Oxidation is losing electrons
- Reduction is gaining electrons:
  - Reducing agent (donates electrons (reducing) other substance)
  - Oxidising agent accepts electrons (oxidising itself)
- E.g.
  - $\text{ZnO} + \text{CO} \rightarrow \text{Zn} + \text{CO}_2$ 
    - $\text{ZnO} \rightarrow \text{Zn}^{2+} + 2\text{e}^-$
    - $\text{CO} + 2\text{e}^- \rightarrow \text{CO}^{2-}$

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