4 Hydrogen and chlorine are reacted together to form hydrogen chloride.

$$H_2(g) + Cl_2(g) \rightarrow 2HCl(g)$$
 $\Delta H = -184 \text{ kJ mol}^{-1}$

(a) Calculate the bond enthalpy for the H-Cl bond using the information in the table below.

| bond | bond enthalpy / kJ mol ⁻¹ |
|-------|---|
| н-н | +436 |
| CI-CI | +243 |

Bonds broken 436

243

= 679

-679

Bonds made 2(HCI)

Bonds Broken + Bonds Made = -184 -679 + 2HCl = -184 2HCl = 495 HCl = 247.5

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